

## Quiz & Worksheet - Limiting Reactants & Excess Reactants

### 1. What is a limiting reactant?

- The reactant that is used up first and prevents more product from being made.
- The reactant that is used up last and prevents more product from being made.
- The reactant that makes the product.
- The substance that is in excess that doesn't get used up as a reactant.

### 2. Say you take a reactant A and calculate the amount of moles of another reactant B required to use up all of A. How do you know which of two reactants is the limiting one?

- You compare the calculated amount of B to the actual amount available. If more is required, then B is the limiting reactant. If less is required, then A is the limiting reactant.
- You compare the calculated amount of B to the actual amount available. If more is required, then A is the limiting reactant. If less is required, then B is the limiting reactant.
- If the calculated value of B is larger than the amount of A, then A is the limiting reactant.
- If the calculated value of B is larger than the amount of A, then B is the limiting reactant.

### 3. If 2.5 moles of zinc react with 6.0 moles of hydrochloric acid in the equation $\text{Zn} + 2\text{HCl} = \text{ZnCl}_2 + \text{H}_2$ , what is the limiting reactant?

- Zn
- HCl
- ZnCl
- ZnH

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# Limiting Reactants Study Guide Answers

**Michal Rosen-Zvi**



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*Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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[Chemical Principles Study Guide/Solutions Manual](#) John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

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