

Limiting Reactant Practice 2.3 - Answer Key

Objective: Use stoichiometry to identify the limiting and excess reactants.

Directions: Answer the following limiting reactant problems. Show your work!!

1. A reaction between potassium and oxygen gas produces potassium oxide, represented by the following balanced chemical equation:



- a. If 3.00 g of potassium and 5.00 g of oxygen are mixed, how much potassium oxide is produced?

$$3.00 \text{ g K} \times \left(\frac{1 \text{ mol K}}{39.10 \text{ g K}} \right) \left(\frac{1 \text{ mol O}_2}{2 \text{ mol K}} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = \boxed{1.21 \text{ g K}_2\text{O}}$$

$$5.00 \text{ g O}_2 \times \left(\frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right) \left(\frac{1 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = 5.00 \text{ g K}_2\text{O}$$

- b. What is the limiting reactant?

Potassium (K)

- c. What is the reactant in excess?

Oxygen (O₂)

2. In a reaction, 12.00 g of aqueous sodium hydroxide reacts with aqueous 22.00 g of copper(II) sulfate.



- a. What mass of sodium sulfate is produced?

$$12.00 \text{ g NaOH} \times \left(\frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{2 \text{ mol NaOH}} \right) \left(\frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = 17.10 \text{ g}$$

$$22.00 \text{ g CuSO}_4 \times \left(\frac{1 \text{ mol CuSO}_4}{159.61 \text{ g CuSO}_4} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol CuSO}_4} \right) \left(\frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = \boxed{19.87 \text{ g}}$$

- b. What is the limiting reactant?

Copper(II) sulfate (CuSO₄)

- c. What is the reactant in excess?

Sodium Hydroxide (NaOH)

Limiting Reactants Study Guide Answer Key

JL Elias



Limiting Reactants Study Guide Answer Key:

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In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Limiting Reactants Study Guide Answer Key**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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