

Chemistry Worksheet: Limiting Reactant Worksheet #1

1. Consider the following reaction: $2 \text{Al} + 6 \text{HBr} \rightarrow 2 \text{AlBr}_3 + 3 \text{H}_2$
- When 3.22 moles of Al reacts with 4.96 moles of HBr, how many moles of H_2 are formed?
 - What is the limiting reactant?
 - For the reactant in excess, how many moles are left over at the end of the reaction?

a) 4.83 mol H_2 and 2.48 mol H_2
b) HBr is the limiting reactant
c) 1.57 mol Al is left

2. Consider the following reaction: $3 \text{Si} + 2 \text{N}_2 \rightarrow \text{Si}_3\text{N}_4$
- When 21.44 moles of Si reacts with 17.62 moles of N_2 , how many moles of Si_3N_4 are formed?
 - What is the limiting reactant?
 - For the reactant in excess, how many moles are left over at the end of the reaction?

a) 7.14 mol Si_3N_4 and 8.81 mol Si_3N_4
b) silicon is the limiting reactant
c) 3.34 mol N_2 are left

3. Consider the following reaction: $2 \text{CuCl}_2 + 4 \text{KI} \rightarrow 2 \text{CuI} + 4 \text{KCl} + \text{I}_2$
- When 0.56 moles of CuCl_2 reacts with 0.64 moles of KI, how many moles of I_2 are formed?
 - What is the limiting reactant?
 - For the reactant in excess, how many moles are left over at the end of the reaction?

a) 0.28 and 0.16 mol I_2
b) KI is the limiting reactant
c) 0.24 mol CuCl_2 are left

4. Consider the following reaction: $4 \text{FeS}_2 + 11 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3 + 8 \text{SO}_2$
- When 26.62 moles of FeS_2 reacts with 5.44 moles of O_2 , how many moles of SO_2 are formed?
 - What is the limiting reactant?
 - For the reactant in excess, how many moles are left over at the end of the reaction?

a) 5.324 and 3.95 mol SO_2
b) O_2
c) 51.27 mol FeS_2 are left

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J Ma



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