CHEMISTRY: LIMITING REAGENT Practice Problems

Please show your work on a separate piece of paper!

Assume standard temperature and pressure, STP conditions for all reactions.

At high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide:

$$Fe_{(0)} + S_{(0)} \rightarrow FeS_{(0)}$$

In one experiment, 7.62 g of Fe are allowed to react with 8.67 g of S. a. What is the limiting reagent, and what is the reactant in excess?

b. Calculate the mass of FeS formed.

 Arcylonitrile, C 3H3N, is the starting material for the production of a kind of synthetic fiber (acrylics) and can be made from propylene, C3H6, by reaction with nitric oxide, NO, as follows:

$$4 C_3H_{6 (g)} + 6 NO_{(g)} \rightarrow 4 C_3H_3N_{(s)} + 6 H_2O_{(l)} + N_{2 (g)}$$

What mass of C₃H₃N can be made when 21.6 g of C₃H₆ react with 21.6 g of nitric oxide?

- Calculate the percent yield for the reaction: P_{4 to} + 6 Cl_{2 tg} → 4 PCl_{3 (b)} if 75.0 g of phosphorus reacts with excess chlorine gas to produce 111.0 g of phosphorus trichloride.
- 4. Formic acid, HCHO2, burns in oxygen to form carbon dioxide and water as follows:

If a 3.15-g sample of formic acid was burned in 2.0 L of oxygen, what volume of carbon dioxide would be produced?

Zinc metal reacts with hydrochloric acid to produce zinc chloride and hydrogen gas.
 a. Balance the following reaction: Zn_(i) + HCl_(aq) → ZnCl_{2 (aq)} + H_{2 (g)}
 b. A 3.50-g sample of zinc metal is allowed to react with 2.50 g of hydrochloric acid.

Complete the following table:

reactants/products	Zn (grams)	HCl (grams)	ZnCl ₂ (grams)	$H_2(L)$
Before reaction				
After reaction	1.26 g			

Limiting Reactant Problems Answer Key

Purshotam Kaushik & Kirti Kaushik

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