

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table

chemistry

LIMITING REACTANTS

1 The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
 If you start with 10 moles of N_2 and 30 moles of H_2 , determine the limiting reactant.

2 Calculate the grams of ammonia produced.

3 How many grams of the excess reactant is produced in a left over?
 To multiply the moles of the limiting reactant by the molar mass of the excess reactant.

4 Calculate moles of product & mass of product.

5 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$
 If you have 100 grams of glucose and 60 grams of oxygen, determine the limiting reactant.

6 How many grams of water (steam) are produced?

7 Calculate moles of product & mass of product.



8 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure iron ore is produced.
 $Mg + 2FeCl_2 \rightarrow MgCl_2 + 2Fe$
 If you start with 10 moles of magnesium and 2 moles of iron(II) chloride, determine the limiting reactant.

9 Calculate the grams of magnesium chloride produced.

10 How many grams of the excess reactant is produced in a left over?

11 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure iron ore is produced.
 $Mg + 2FeCl_2 \rightarrow MgCl_2 + 2Fe$
 If you have with 100 grams of magnesium and 60 grams of iron(II) chloride, determine the limiting reactant.

12 How much carbon dioxide is produced?

13 Calculate moles of product & mass of product.

12 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$
 If you have 100 grams of glucose and 60 grams of oxygen, determine the limiting reactant.

13 Calculate moles of product & mass of product.

14 How many grams of the excess reactant is produced in a left over?

15 How many grams of the excess reactant is produced in a left over?

16 How many grams of the excess reactant is produced in a left over?

17 How many grams of the excess reactant is produced in a left over?

18 How many grams of the excess reactant is produced in a left over?

ANSWER KEY INCLUDED

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AW Rasmussen



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