

## Limiting Reactants Worksheet KEY

Answer the following questions, using stoichiometry to convert between moles of one substance and moles of another substance within a reaction and using molar masses to convert between grams of a substance and moles of that substance.

Given the following reaction:  $\text{C}_5\text{H}_{12} + 8 \text{O}_2 \rightarrow 5 \text{CO}_2 + 6 \text{H}_2\text{O}$

1. If 5.25 grams of pentane ( $\text{C}_5\text{H}_{12}$ ) are completely reacted with oxygen ( $\text{O}_2$ ), how many grams of carbon dioxide ( $\text{CO}_2$ ) will be produced?

$$5.25 \text{ g C}_5\text{H}_{12} \times \frac{1 \text{ mol C}_5\text{H}_{12}}{72.151 \text{ g C}_5\text{H}_{12}} = 0.07276 \text{ mol C}_5\text{H}_{12}$$

$$0.07276 \text{ mol C}_5\text{H}_{12} \times \frac{5 \text{ mol CO}_2}{1 \text{ mol C}_5\text{H}_{12}} = 0.3638 \text{ mol CO}_2$$

$$0.3638 \text{ mol CO}_2 \times \frac{44.009 \text{ g CO}_2}{1 \text{ mol CO}_2} = 16.0 \text{ g CO}_2$$

2. If 9.75 grams of oxygen ( $\text{O}_2$ ) are completely reacted with pentane ( $\text{C}_5\text{H}_{12}$ ), how many grams of carbon dioxide ( $\text{CO}_2$ ) will be produced?

$$9.75 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{31.998 \text{ g O}_2} = 0.3047 \text{ mol O}_2$$

$$0.3047 \text{ mol O}_2 \times \frac{5 \text{ mol CO}_2}{8 \text{ mol O}_2} = 0.1904 \text{ mol CO}_2$$

$$0.1904 \text{ mol CO}_2 \times \frac{44.009 \text{ g CO}_2}{1 \text{ mol CO}_2} = 8.38 \text{ g CO}_2$$

3. Based on your answers to question #1 and 2, if 5.25 grams of pentane are reacted with 9.75 grams of oxygen, what is the limiting reactant and how many grams of  $\text{CO}_2$  actually form?

**Limiting reactant:  $\text{O}_2$**

**8.38 g  $\text{CO}_2$  actually form**

# Limiting Reactant Answer Key

**Purshotam Kaushik & Kirti Kaushik**

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**Introduction to General, Organic, and Biochemistry** Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15

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**Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28. The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

**Chemistry** Greg Curran, 2011. Covers all the topics in a typical one year high school chemistry curriculum.

Microbiology (Questions and Answers), 5e Purshotam Kaushik & Kirti Kaushik, 2022. Microbiology is an engaging textbook presenting a balanced and comprehensive account of major areas of microbiology in the form of questions and answers. This question answer approach to present complex topics and theories of microbiology regarding cellular and non cellular microorganisms, microbial genetics, and molecular biology in higher plants and animals makes the subject interesting and easily comprehensible for the students.

**NEET Chemistry 1500+ MCQs** Disha Experts, 2019-12-24.

*Mapping College Chemistry* Stephen DeMeo, 2019-03-01. This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course. The book incorporates a chemistry problem solving plan, one that uses an innovative graphic organizer strategy. The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry. Topical problem solving will focus on limiting reactant, stoichiometry, identifying types of chemical reactions, equilibrium, acid base equilibria, and electrochemistry. Why would this resource be of interest to chemistry students? To be successful to get into a well known college, medical school, physical therapy, or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course. To make matters worse, many college professors feel that only a few students should get A grades, and therefore they give difficult exams that many students fail. This is the weeding out process that every pre health student is apprehensive about. To succeed in this competitive environment entails not just studying harder or longer; it means re organizing textbook content so that it is meaningful to the student. This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve

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