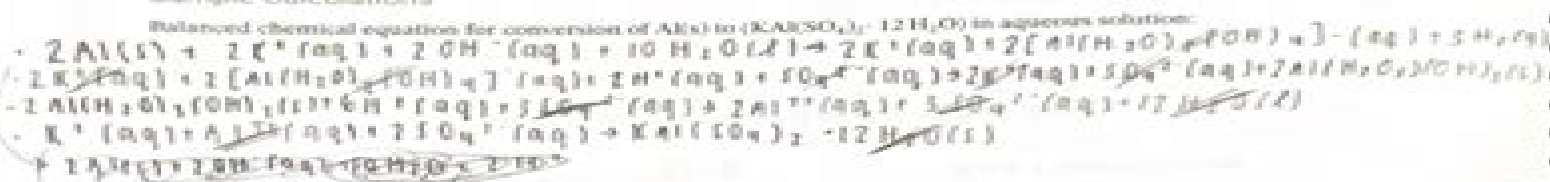


Sample Calculations

Balanced chemical equation for conversion of Al(s) to $(\text{KAl}(\text{SO}_4)_2 \cdot 12 \text{H}_2\text{O})$ in aqueous solution:

Theoretical Yield:

• Moles alum • MW alum (g/mol) = grams of alum

$$0.83 \text{ g Al} \left(\frac{1 \text{ mol Al}}{26.982 \text{ g Al}} \right) \left(\frac{2 \text{ mol KAl}(\text{SO}_4)_2 \cdot 12 \text{H}_2\text{O}}{2 \text{ mol Al}} \right) \left(\frac{474.39 \text{ g alum}}{1 \text{ mol alum}} \right) = 14.59 \text{ g}$$

Percent Yield:

$$\frac{\text{actual}}{\text{theo}} \times 100$$

$$\frac{11.56}{14.59} \times 100 = 79.23\%$$



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Full Committee Briefing on the Recently Proposed Chemical Arms Treaty United States. Congress. House. Committee on Armed Services, 1984

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