

SOLUTIONS

The Laws of Thermodynamics

- **0th Law**
Introduces the thermodynamic intensive variable of temperature (T)
- **1st Law**
Conservation and conversion of energy
Defines extensive thermodynamic state variable of internal energy (E)
 $dE = dq + dw$
- **2nd Law**
Defines the extensive thermodynamic state variable of entropy (S)
 $dS_{univ} \geq 0$
- **3rd Law**
For systems in internal equilibrium, sets the zero of entropy at the minimum in temperature (0K) and at the minimum in internal energy

The Three EoS Equations

$$dS = \frac{1}{T}dT + \frac{P}{T}dV$$

$$dE = TdS - PdV$$

$$dH = TdS + VdP$$

Partial Derivatives of Thermodynamic Properties

$$\left(\frac{\partial E}{\partial S}\right)_V = T, \quad \left(\frac{\partial E}{\partial V}\right)_S = -P$$

$$\left(\frac{\partial H}{\partial S}\right)_P = T, \quad \left(\frac{\partial H}{\partial P}\right)_S = V$$

$$\left(\frac{\partial A}{\partial T}\right)_V = -S, \quad \left(\frac{\partial A}{\partial V}\right)_T = P$$

$$\left(\frac{\partial G}{\partial T}\right)_P = -S, \quad \left(\frac{\partial G}{\partial P}\right)_T = V$$

Maxwell Relations for Single Component

$$\left(\frac{\partial T}{\partial V}\right)_S = -\left(\frac{\partial P}{\partial S}\right)_V$$

$$\left(\frac{\partial T}{\partial P}\right)_S = \left(\frac{\partial V}{\partial S}\right)_P$$

$$\left(\frac{\partial S}{\partial V}\right)_T = \left(\frac{\partial P}{\partial T}\right)_V$$

$$\left(\frac{\partial S}{\partial P}\right)_T = -\left(\frac{\partial V}{\partial T}\right)_P$$

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Gaskell Solution Manual

Rachel S Tattersall



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