

## Questions and Problems

**Q1** According to your graph, what is the relationship between the pressure and the volume of a gas?

\* As the volume of the gas increases the pressure of the gas decreases. They have inverse proportional relationship.

**Q2** On your graph, what is the volume of the gas at a pressure of 760. mmHg?

\* After plotting the data to excel, the graph produced a linear equation  $y = -28.873x + 1539.3$  where y is the pressure and x is the volume of gas.

We need to find x at 760 mmHg.

Hence,

$$760 \text{ mmHg} = -28.873x + 1539.8 \Rightarrow x = 26.99 \text{ mL}$$

Therefore, at 760 mmHg the volume of the gas is 26.99 mL

**Q3** On your graph, what is the pressure of the gas when the volume is 30.0 mL?

\*  $x = 30.0 \text{ mL}$ ,  $y = ?$

$$y = -28.873 (30.0) + 1539.3$$

$$y = 673.1 \text{ mmHg}$$

Therefore, when the volume of the gas is 30.0 mL the pressure is 673.1 mmHg

**Q4** A sample of helium has a volume of 325 mL and a pressure of 655 mmHg. What will be the pressure, in mmHg, if the sample of helium is compressed to 125 mL. (T, n constant)?

$$P_1 V_1 = P_2 V_2$$

$$P_2 = (P_1 V_1) \div V_2 = (655 \text{ mmHg} \times 325 \text{ mL}) \div 125 \text{ mL} = 1,703 \text{ mmHg}$$

# Gas Laws Section Review Answers By Pearson

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In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Gas Laws Section Review Answers By Pearson**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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