

## Chemistry Laboratory 11B - The Iodine-Clock Reaction Shawn Yoon

### Purpose:

The purposes of this laboratory were to:

- observe and record the effect of changing the concentrations of a reactant on the rate of reactions, and to
- observe and record the effect of changing the temperature of a system on the rate of a reaction.

### Materials and Methods:

The materials and methods for this laboratory followed the procedures laid out in Experiment 11B on pages 149-170 of *Essential Experiments for Chemistry* by Morrison and Boyd with the following exceptions:

- only two beakers were used to make water baths instead of four,
- volumetric pipettes were used instead of dropping pipettes,
- 25 mm + 100 mm test tubes were used instead of 15 mm + 175 mm ones,
- four test tubes were used instead of eight, and
- solutions A and B were mixed by shaking the test tube slightly instead of pouring between both test tubes.

### Observations:

#### Part I:

First, approximately 50 mL of solution A and 50 mL of solution B were obtained from the front of the laboratory in 100 mL beakers. According to the instructor, solution A contained 0.002M  $\text{KIO}_3$  and solution B contained 0.0005M  $\text{NaHSO}_3$ ,  $\text{H}_2\text{SO}_4$ , and starch. Solution A appeared very clear, its appearance no different than water while solution B appeared very cloudy, due to the starch in solution.

Then, four 25 mm + 100 mm test tubes were obtained from the laboratory and placed in a test tube rack. Using a volumetric pipette, 10 mL of solution A was transferred into a graduated

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**Richard Kidder Meade, Harry  
McCormack, Laurance T.  
Clark, Alexander G. Sclater, Lloyd  
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