

Empirical and Molecular Formula Practice KEY

Empirical Formula

1. 36.5% Na, 25.4% S, and 38.1% O forms this compound.

$$\text{Na} - 36.5/23 = 1.56/0.8 = 2$$

$$\text{S} - 25.4/32 = 0.8/0.8 = 1$$

$$\text{O} - 38.1/16 = 2.4/0.8 = 3$$



2. A compound made of 72.2% Magnesium and 27.8% nitrogen by mass.

$$\text{Mg} - 72.2/24.3 = 3/2 = 1.5(2) = 3$$

$$\text{N} - 27.8/14 = 2/2 = 1(2) = 2$$



3. A chromium compound with 68.4% Cr and 31.6% O.

$$\text{Cr} - 68.4/52 = 1.315/1.315 = 1(2) = 2$$

$$\text{O} - 31.6/16 = 1.975/1.315 = 1.5(2) = 3$$



4. A compound with the composition of 24% C, 7% H, 38% F, and 31% P.

$$\text{C} - 24/12 = 2/1 = 2$$

$$\text{H} - 7/1 = 7/1 = 7$$

$$\text{F} - 38/19 = 2/1 = 2$$

$$\text{P} - 31/31 = 1/1 = 1$$



5. 63.5% silver, 8.2% nitrogen, and 28.3% oxygen make up this compound.

$$\text{Ag} - 63.5/107.87 = 0.589/0.585 = 1$$

$$\text{N} - 8.2/14.01 = 0.585/0.585 = 1$$

$$\text{O} - 28.3/16 = 1.769/0.585 = 3$$



6. 8.81g C and 91.2g Chlorine forms this compound.

$$\text{C} - 8.81/12 = 0.73/0.73 = 1(2) = 2$$

$$\text{Cl} - 91.2/35 = 2.61/0.73 = 3.5(2) = 7$$



7. A compound in a soft drink that has 0.31g hydrogen, 3.2g phosphorus, and 6.5g oxygen.

$$\text{H} - 0.31/1 = 0.31/.1 = 3.1$$

$$\text{P} - 3.2/31 = 0.10/.1 = 1$$

$$\text{O} - 6.5/16 = 0.40/.1 = 4$$



Empirical And Molecular Formulas Answers

Eileen Ramsden



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Basics for Chemistry David A. Ucko, 2013-09-24 Basics of Chemistry provides the tools needed in the study of General Chemistry such as problem solving skills calculation methods and the language and basic concepts of chemistry The book is designed to meet the specific needs of underprepared students Concepts are presented only as they are needed and developed from the simple to the complex The text is divided into 18 chapters each covering some particular aspect of chemistry such as matter energy and measurement the properties of atoms description of chemical bonding study of chemical change and nuclear and organic chemistry Undergraduate students will find the book as a very valuable academic material

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