

1. An unknown compound was found to have a percent composition as follows:  
47.0 % potassium, 14.5 % carbon, and 38.5 % oxygen. What is its empirical formula?

$$\begin{aligned} K &= 47.0 \left( \frac{1}{39} \right) = 1.205 / 1.2 = 1 \\ C &= 14.5 \left( \frac{1}{12} \right) = 1.208 / 1.2 = 1 \\ O &= 38.5 \left( \frac{1}{16} \right) = 2.406 / 1.2 = 2 \end{aligned} \quad \boxed{KCO_2}$$

2. A sample of iodine chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the iodine compound?

$$\begin{aligned} I &= .2596 \left( \frac{1}{127} \right) = .00204 / .0012 = 1 \\ Cl &= .2404 \left( \frac{1}{35.5} \right) = .0067 / .0012 = 3 \end{aligned} \quad \boxed{ICl_3}$$

3. Nitrogen and oxygen form an extensive series of oxides with the general formula  $N_xO_y$ . One of them is a blue solid that comes apart, reversibly, in the gas phase. It contains 36.84% N. What is the empirical formula of this oxide?

$$\begin{aligned} N &= 36.84 \left( \frac{1}{14} \right) = 2.631 / 2.63 = 1 + 2 = 2 \\ O &= 63.16 \left( \frac{1}{16} \right) = 3.9475 / 2.63 = 1.5 + 2 = 3 \end{aligned} \quad \boxed{N_2O_3}$$

4. Rubbing alcohol was found to contain 60.0 % carbon, 13.4 % hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of rubbing alcohol?

$$\begin{aligned} C &= 60 \left( \frac{1}{12} \right) = 5 \text{ Atoms} = 5 \\ H &= 13.4 \left( \frac{1}{1} \right) = 13.4 / 1.4 = 10 \\ O &= 26.6 \left( \frac{1}{16} \right) = 1.66 / 1.6 = 1 \end{aligned} \quad \boxed{C_5H_{10}O}$$

5. What is the empirical formula of a compound that contains 0.783g of Carbon, 0.186g of Hydrogen and 0.521g of Oxygen?

$$\begin{aligned} C &= .783 \left( \frac{1}{12} \right) = .06525 = 2 \\ H &= .186 \left( \frac{1}{1} \right) = .186 = 6 \\ O &= .521 \left( \frac{1}{16} \right) = .03256 = 1 \end{aligned} \quad \boxed{C_2H_6O}$$

6. What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\begin{aligned} Au &= 89.14 \left( \frac{1}{197} \right) = .4525 / .3 = 1 + 2 = 2 \\ O &= 10.8 \left( \frac{1}{16} \right) = .675 / .3 = 1.5 + 2 = 3 \end{aligned} \quad \boxed{Au_2O_3}$$

# Instructional Fair Determining Empirical Formulas

**EW Minium**



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**Empirical Direction in Design and Analysis** Norman H. Anderson, 2001-07-01 The goal of Norman H Anderson s new book is to help students develop skills of scientific inference To accomplish this he organized the book around the Experimental Pyramid six levels that represent a hierarchy of considerations in empirical investigation conceptual framework phenomena behavior measurement design and statistical inference To facilitate conceptual and empirical understanding Anderson de emphasizes computational formulas and null hypothesis testing Other features include emphasis on visual inspection as a basic skill in experimental analysis to help students develop an intuitive appreciation of data patterns exercises that emphasize development of conceptual and empirical application of methods of design and analysis and de emphasize formulas and calculations and heavier emphasis on confidence intervals than significance tests The book is intended for use in graduate level experimental design research methods or statistics courses in psychology education and other applied social sciences as well as a professional resource for active researchers The first 12 chapters present the core concepts graduate students must understand The next nine chapters serve as a reference handbook by focusing on specialized topics with a minimum of technicalities

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*Machinery* Fred Herbert Colvin, Lester Gray French, 1922

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1. Understanding the eBook Instructional Fair Determining Empirical Formulas
  - The Rise of Digital Reading Instructional Fair Determining Empirical Formulas
  - Advantages of eBooks Over Traditional Books
2. Identifying Instructional Fair Determining Empirical Formulas
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Instructional Fair Determining Empirical Formulas
  - User-Friendly Interface
4. Exploring eBook Recommendations from Instructional Fair Determining Empirical Formulas
  - Personalized Recommendations
  - Instructional Fair Determining Empirical Formulas User Reviews and Ratings
  - Instructional Fair Determining Empirical Formulas and Bestseller Lists
5. Accessing Instructional Fair Determining Empirical Formulas Free and Paid eBooks
  - Instructional Fair Determining Empirical Formulas Public Domain eBooks
  - Instructional Fair Determining Empirical Formulas eBook Subscription Services
  - Instructional Fair Determining Empirical Formulas Budget-Friendly Options

6. Navigating Instructional Fair Determining Empirical Formulas eBook Formats
  - ePub, PDF, MOBI, and More
  - Instructional Fair Determining Empirical Formulas Compatibility with Devices
  - Instructional Fair Determining Empirical Formulas Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Instructional Fair Determining Empirical Formulas
  - Highlighting and Note-Taking Instructional Fair Determining Empirical Formulas
  - Interactive Elements Instructional Fair Determining Empirical Formulas
8. Staying Engaged with Instructional Fair Determining Empirical Formulas
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Instructional Fair Determining Empirical Formulas
9. Balancing eBooks and Physical Books Instructional Fair Determining Empirical Formulas
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Instructional Fair Determining Empirical Formulas
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Instructional Fair Determining Empirical Formulas
  - Setting Reading Goals Instructional Fair Determining Empirical Formulas
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Instructional Fair Determining Empirical Formulas
  - Fact-Checking eBook Content of Instructional Fair Determining Empirical Formulas
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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