

- 2 -

1. Which of the following processes provide evidence of the particulate nature of matter?

I. Capillarity
II. Diffusion
III. Filtration
IV. Osmosis

(A) I and III only
(B) I and IV only
(C) II and III only
(D) II and IV only

2. 'Mass number' is the number of

(A) neutrons plus protons
(B) neutrons minus protons
(C) electrons plus neutrons
(D) electrons plus protons

3. Which of the following elements has 7 electrons in its outer shell?

(A) Hydrogen
(B) Oxygen
(C) Nitrogen
(D) Chlorine

4. An atom that has 20 electrons when it ionizes has the same electronic configuration as

(A) argon
(B) calcium
(C) neon
(D) sodium

5. $^{23}_{11}\text{Na}$ and $^{24}_{11}\text{Na}$ are isotopes because they

(A) conduct electricity
(B) react vigorously with water
(C) have the same number of protons
(D) have the same relative atomic mass

6. Sulfur and oxygen are in the same group of the periodic table because

(A) they can react with each other
(B) they can form covalent compounds
(C) they have the same number of electrons in their outer shell
(D) the atomic number of sulfur is 16 and the relative atomic mass of oxygen is 16

Item 7 refers to the following quantities of atoms.

I. 12 g of carbon
II. 16 g of sulfur
III. 23 g of sodium
IV. 12 g of magnesium

(Relative mass number: C = 12; S = 32; Mg = 24; Na = 23)

7. Which of the following pairs represents equal quantities of atoms?

(A) I and II only
(B) I and IV only
(C) II and III only
(D) II and IV only

8. An ion with a single negative charge may be converted into a neutral atom by

(A) losing a neutron
(B) gaining a proton
(C) losing an electron
(D) gaining an electron

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Ron Holt



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