

Chapter 10: Chemical Energy and Thermochemistry

1. What makes exothermic reactions safe to handle?
2. Account for the combustion of fuels (exothermic reactions).
3. The amount of energy required to raise the temperature of a kilogram of a substance by 1 kelvin is called its specific heat.
4. As the chemical energy of matter increases, its potential energy stored in a system also increases.
5. Thermodynamics can be defined as the study of energy and its transformations.
6. The amount of chemical energy that is released when a substance (fuel) is oxidized is called its heat of combustion.
7. The amount of energy for a thermodynamic system is called its internal energy.
8. What are the two basic thermodynamic laws? How do they relate to energy and its transformations?
9. Explain the first law of thermodynamics (conservation of energy).
10. Explain the second law of thermodynamics (entropy).
11. As the temperature of a system increases, the average kinetic energy of the particles in the system increases.
12. A system's internal energy is the sum of the kinetic and potential energies of its particles.
13. Explain the third law of thermodynamics (entropy).
14. Explain the concept of a good conductor of heat.
15. Account for the energy changes that occur when a substance is heated or cooled.
16. The amount of heat that is released when a substance is oxidized is called its heat of combustion.
17. Explain the concept of a good conductor of heat.
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Energy And Heat Study Guide Answers

Wayne R. Hedrick, Paul R. Wagner



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