

METALS

A **metal** is an element which has the ability to lose electron(s) and become positively charged. There are about 105 different elements. Of these 84 are metals and 21 are non metals.

GENERAL PROPERTIES

Metals generally have the following properties.

- i. Have high melting and boiling points
- ii. Are good conductors of heat and electricity
- iii. Have a high density – metal atoms are packed very closely
- iv. Are solids at rtp, except for mercury which is a liquid.
- v. Are *Malleable* i.e. can be beaten into sheets
- vi. Are *Ductile* i.e. can be made into wires.
- vii. Are *Sonorous* i.e. produce sound when beaten
- viii. Are *Lustrous* i.e. can be polished or have a shining surface.
- ix. They are *strong* under tension and compression. That means they can withstand stretching and crushing without breaking
- x. React with oxygen to form basic oxides
- xi. Form positively charged ions (Cations) by losing electrons.

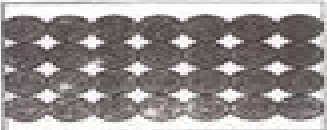
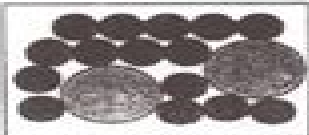
WHY METALS ARE USED AS ALLOYS

Pure metals are usually too soft and weak for most uses. A gold chain is never made of pure gold, as the chain would be too soft and would break easily. Instead, it is alloyed with other elements to make it more shiny and stronger.

Metals are mainly used as alloys for the following reasons

- To improve on their properties.
- To improve on their appearance.
- To make metals to be less brittle.
- To Make metals strong and harder.

A comparison between pure metals and alloys is shown below.

Pure metal	Alloy
Are usually Soft and can easily bent	Are Strong and hard, not easily bent
Have higher melting points.	Have lower melting points than the melting points of their components
Atoms can slide over each other easily as atoms are of the same size  <i>Diagrammatic representation of a pure metal</i>	Atoms cannot slide over each other easily as foreign atoms are of different size and disrupt the orderly arrangement of the metal atoms  <i>Diagrammatic representation of an alloy</i>

Metals Study Guide Answers

Clifford Lane

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Table of Contents Metals Study Guide Answers

1. Understanding the eBook Metals Study Guide Answers
 - The Rise of Digital Reading Metals Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Metals Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Metals Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Metals Study Guide Answers
 - Personalized Recommendations
 - Metals Study Guide Answers User Reviews and Ratings
 - Metals Study Guide Answers and Bestseller Lists
5. Accessing Metals Study Guide Answers Free and Paid eBooks
 - Metals Study Guide Answers Public Domain eBooks
 - Metals Study Guide Answers eBook Subscription Services
 - Metals Study Guide Answers Budget-Friendly Options
6. Navigating Metals Study Guide Answers eBook Formats

- ePub, PDF, MOBI, and More
- Metals Study Guide Answers Compatibility with Devices
- Metals Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Metals Study Guide Answers
 - Highlighting and Note-Taking Metals Study Guide Answers
 - Interactive Elements Metals Study Guide Answers
- 8. Staying Engaged with Metals Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Metals Study Guide Answers
- 9. Balancing eBooks and Physical Books Metals Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Metals Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Metals Study Guide Answers
 - Setting Reading Goals Metals Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Metals Study Guide Answers
 - Fact-Checking eBook Content of Metals Study Guide Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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