Chapter 4 Atomic Structure

Section 4.2 The Structure of an Atom (pages 108–112)

This section compares the properties of three subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.

Reading Strategy (page 108)

Monitoring Your Understanding Before you read, list in the table shown what you know about atoms and what you would like to learn. After you read, list what you have learned. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook.

What I Know	What I Would	What I Have
About Atoms	Like to Learn	Learned
Most students will know that atoms are the "building blocks" of matter, and some may know that atoms contain subatomic particles.	Based on the title of the section, students may say that they want to learn more about the structure of atoms.	

Properties of Subatomic Particles (pages 108-109)

1. What are three subatomic particles?

a. Protons b. Electrons c. Neutrons

Circle the letter that identifies a subatomic particle with a positive charge.

a. nucleus

(b) proton

c. neutron

d. electron

Why did Chadwick conclude that the particles produced by his experiment were neutral in charge? A charged object did not deflect the paths of the particles.

Comparing Subatomic Particles (pages 109-110)

Circle the letters of properties that vary among subatomic particles.

a. color

(b) mass

c. charge

d location in the atom

Circle the letter of the expression that accurately compares the masses of neutrons and protons.

a mass of 1 neutron = mass of 1 proton

b. mass of 2000 neutrons = mass of 1 proton

c. mass of 1 electron = mass of 1 proton

d. mass of 1 neutron = mass of 1 electron

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Clifford Lane

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