

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

### Midterm Exam Study Guide

#### d. Isotopes

- i. Isotopes differ in their \_\_\_\_\_ and number of \_\_\_\_\_

#### e. Average atomic mass Formula:

#### f. Ions

- i. \_\_\_\_\_ have lost electrons and have a \_\_\_\_\_ charge
- ii. \_\_\_\_\_ have gained electrons and have a \_\_\_\_\_ charge

#### g. Group Numbers be able to locate them

- i. Alkali Metals
- ii. Alkaline earth Metals
- iii. Transition Metals
- iv. Inner Transition Metals
- v. Halogens
- vi. Noble Gases

#### h. Periodic Trends Give your notes for reference (the charts for each of these)

- i. Atomic Radius/Size
- ii. Ionic Radius/Size
- iii. Ionization Energy
- iv. Electronegativity

#### i. Valence Electrons be able to identify # of valence electrons and charge for each group

- i. Lewis-Dot structures

### F. Nuclear Decay and Half Life

- a.  ${}^4_2\text{He}$  or  ${}^4_2\text{He}$  is called an \_\_\_\_\_ particle which has the weakest radiation.
- b.  ${}^0_{-1}\text{e}$  or  ${}^0_{-1}\beta$  is called an \_\_\_\_\_ particle which has moderate radiation and is an electron.
- c.  ${}^4_{-2}\text{He}$  is called an \_\_\_\_\_ particle which has the strongest radiation and highest energy.
- d. Half-life
  - i. Definition: \_\_\_\_\_
  - ii. Know how to read a graph and calculate half-life

### d. Electron Configurations

- a. Be able to draw them/identify them from another student
- b. Hund's Rule
- c. Pauli Exclusion Principle
- d. Aufbau Principle

### e. Ionic Compounds

- a. Ionic compounds are made of \_\_\_\_\_ + \_\_\_\_\_
- b. Their overall charge is \_\_\_\_\_
- c. Know how to use Lewis-Dot Structures to draw bonds between cation and anion

# Mcps Chemistry Semester Review Guide Answers

**VM Jensen**



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