

5 • Chemical Bonding**LEWIS STRUCTURES**Indicate the # of VALENCE electrons for each species. Write the correct Lewis electron dot structure for each.

F # of electrons = <u>7</u> $\cdot \ddot{\text{F}} \cdot$	O # of electrons = <u>6</u> $\begin{array}{c} \cdot \ddot{\text{O}} \cdot \\ \text{or} \\ \cdot \ddot{\text{O}} \cdot \end{array}$	K # of electrons = <u>1</u> $\text{K} \cdot$	Al # of electrons = <u>3</u> $\cdot \text{Al} \cdot$
F⁻ <u>7+1</u> # of electrons = <u>8</u> $[\cdot \ddot{\text{F}}:]^{-}$	O²⁻ <u>6+2</u> # of electrons = <u>8</u> $[\cdot \ddot{\text{O}}:]^{2-}$	K⁺ <u>1-1</u> # of electrons = <u>0</u> $[\text{K}]^{+}$	Al³⁺ <u>3-3</u> # of electrons = <u>0</u> $[\text{Al}]^{3+}$
F₂ <u>7+7</u> # of electrons = <u>14</u> $\cdot \ddot{\text{F}} : \ddot{\text{F}} \cdot$	H₂ <u>1+1</u> # of electrons = <u>2</u> $\text{H} : \text{H}$	HF <u>1+7</u> # of electrons = <u>8</u> $\text{H} : \ddot{\text{F}} \cdot$	NH₃ <u>5+3</u> # of electrons = <u>8</u> $\begin{array}{c} \text{H} : \ddot{\text{N}} : \text{H} \\ \\ \text{H} \end{array}$
CH₄ <u>4+4</u> # of electrons = <u>8</u> $\begin{array}{c} \text{H} \\ \\ \text{H} : \ddot{\text{C}} : \text{H} \\ \\ \text{H} \end{array}$	NF₃ <u>5+21</u> # of electrons = <u>26</u> $\begin{array}{c} \cdot \ddot{\text{F}} : \ddot{\text{N}} : \ddot{\text{F}} \cdot \\ \\ \cdot \ddot{\text{F}} \cdot \end{array}$	SiF₄ <u>4+28</u> # of electrons = <u>32</u> $\begin{array}{c} \cdot \ddot{\text{F}} \cdot \\ \\ \cdot \ddot{\text{F}} : \ddot{\text{Si}} : \ddot{\text{F}} \cdot \\ \\ \cdot \ddot{\text{F}} \cdot \end{array}$	C₂H₆ <u>8+6</u> # of electrons = <u>14</u> $\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H} : \ddot{\text{C}} : \ddot{\text{C}} : \text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array}$
MgH₂ <u>2+2</u> # of electrons = <u>4</u> $\text{H} : \text{Mg} : \text{H}$	LiH <u>1+1</u> # of electrons = <u>2</u> $\text{Li} : \text{H}$	AlH₃ <u>3+3</u> # of electrons = <u>6</u> $\begin{array}{c} \text{H} : \text{Al} : \text{H} \\ \\ \text{H} \end{array}$	BH₃ <u>3+3</u> # of electrons = <u>6</u> $\begin{array}{c} \text{H} : \text{B} : \text{H} \\ \\ \text{H} \end{array}$

Lewis Dot Structures Of Atoms And Ions Answer Key

**Dr Vijay Sarda, Dr A C Handa, Dr K K
Arora**



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Regents Chemistry--Physical Setting Power Pack Revised Edition Barron's Educational Series,Albert S. Tarendash,2021-01-05 Barron s two book Regents Chemistry Power Pack provides comprehensive review actual administered exams and practice questions to help students prepare for the Chemistry Regents exam This edition includes Regents Exams and Answers Chemistry Eight actual administered Regents Chemistry exams so students can get familiar with the test Thorough explanations for all answers Self analysis charts to help identify strengths and weaknesses Test taking techniques and strategies A detailed outline of all major topics tested on this exam A glossary of important terms to know for test day Let s Review Regents Chemistry Extensive review of all topics on the test Extra practice questions with answers A detailed introduction to the Regents Chemistry course and exam One actual recently released Regents Chemistry exam with an answer key

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Basic Chemistry Concepts and Exercises John Kenkel,2011-07-08 Chemistry can be a daunting subject for the uninitiated and all too often introductory textbooks do little to make students feel at ease with the complex subject matter Basic Chemistry Concepts and Exercises brings the wisdom of John Kenkel s more than 35 years of teaching experience to communicate the fundamentals of chemistry in a practical down to earth manner Using conversational language and logically assembled graphics the book concisely introduces each topic without overwhelming students with unnecessary detail Example problems and end of chapter questions emphasize repetition of concepts preparing students to become adept at the basics before they progress to an advanced general chemistry course

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Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Darius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

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