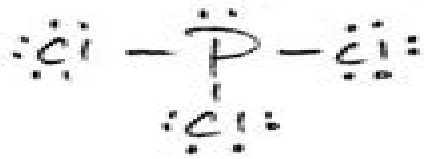


N_2H_4 $14e^-$ $\star$  <i>dinitrogen tetrahydride</i>	$\text{Phosphorus trichloride}$ $5 + 3 \times 7 = 26$ 
NF_3 $5 + 3(7) = 26e^-$  <i>Nitrogen trifluoride</i>	$\text{Tricarbon octahydride}$ (I know, but you can figure it out)  $3(4) + 8 = 20$
Carbon nitride $5 + 4 = 9e^-$ $\star$  <i>"Bot Mr. Jones! That's 10e-! Vep. And it's one of your IONS!"</i>	C_3H_4 $3(4) + 4 = 16e^-$  <i>Tricarbon tetrahydride</i>
$\text{Selenium dihydride}$ $6 + 2 = 8e^-$ 	$\text{Chlorine trifluoride}$ $\star$  <i>expanded octet</i>

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Lewis Structure Worksheet Guided Practice Answers

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