12. Urea, (NH<sub>2</sub>)<sub>2</sub>CO, is an important fertilizer that is manufactured by the following reaction.

$$2NH_3(g) + CO_2(g) \rightarrow (NH_2)_2CO(s) + H_2O(g)$$

What volume of NH<sub>3</sub> at STP will be needed to produce  $8.50 \times 10^4$  kg of urea if there is an 89.5% yield in the process?

13. An obsolete method of generating oxygen in the laboratory involves the decomposition of barium peroxide by the following equation.

$$2\text{BaO}_2(s) \rightarrow 2\text{BaO}(s) + \text{O}_2(g)$$

What mass of BaO<sub>2</sub> reacted if 265 mL of O<sub>2</sub> is collected by water displacement. at 0.975 atm and 10.°C?

14. It is possible to generate chlorine gas by dripping concentrated HCl solution onto solid potassium permanganate according to the following equation.  $2KMnO_4(s) + 16HCl(aq) \rightarrow 2KCl(aq) + 2MnCl_2(aq) + 8H_2O(l) + 5Cl_2(g)$ If excess HCl is dripped onto 15.0 g of KMnO<sub>4</sub>, what volume of Cl<sub>2</sub> will be

produced? The Cl<sub>2</sub> is measured at 15°C and 0.959 atm. Ammonia can be oxidized in the presence of a platinum catalyst according to

the following equation.  $4NH_3(g) + 5O_2(g) \rightarrow 4NO(g) + 6H_2O(l)$ 

The NO that is produced reacts almost immediately with additional oxygen according to the following equation.

$$2NO(g) + O_2(g) \rightarrow 2NO_2(g)$$

If 35.0 kL of oxygen at STP react in the first reaction, what volume of NH<sub>3</sub> at STP reacts with it? What volume of NO<sub>2</sub> at STP will be formed in the second reaction, assuming there is excess oxygen that was not used up in the first reaction?

16. Oxygen can be generated in the laboratory by heating potassium chlorate. The reaction is represented by the following equation.

$$2\text{KClO}_3(s) \rightarrow 2\text{KCl}(s) + 3\text{O}_2(g)$$

What mass of KClO<sub>3</sub> must be used in order to generate 5.00 L of O<sub>2</sub>, measured at STP?

17. One of the reactions in the Solvay process is used to make sodium hydrogen carbonate. It occurs when carbon dioxide and ammonia are passed through concentrated salt brine. The following equation represents the reaction.

 $NaCl(aq) + H_2O(l) + CO_2(g) + NH_3(g) \rightarrow NaHCO_3(s) + NH_4Cl(aq)$ 

- a. What volume of NH<sub>3</sub> at 25°C and 1.00 atm pressure will be required if 38 000 L of CO2, measured under the same conditions, react to form NaHCO<sub>3</sub>?
- b. What mass of NaHCO<sub>3</sub> can be formed when the gases in (a) react with NaCl?

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**Annelies Wilder-Smith** 

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- 1. Understanding the eBook Holt Chemistry Stoichiometry Problem Solving Answers
  - The Rise of Digital Reading Holt Chemistry Stoichiometry Problem Solving Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Holt Chemistry Stoichiometry Problem Solving Answers
  - Exploring Different Genres
  - $\circ\,$  Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Holt Chemistry Stoichiometry Problem Solving Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Holt Chemistry Stoichiometry Problem Solving Answers
  - Personalized Recommendations
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- Holt Chemistry Stoichiometry Problem Solving Answers and Bestseller Lists
- 5. Accessing Holt Chemistry Stoichiometry Problem Solving Answers Free and Paid eBooks
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  - Holt Chemistry Stoichiometry Problem Solving Answers Budget-Friendly Options
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  - Adjustable Fonts and Text Sizes of Holt Chemistry Stoichiometry Problem Solving Answers
  - Highlighting and Note-Taking Holt Chemistry Stoichiometry Problem Solving Answers
  - Interactive Elements Holt Chemistry Stoichiometry Problem Solving Answers
- 8. Staying Engaged with Holt Chemistry Stoichiometry Problem Solving Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Holt Chemistry Stoichiometry Problem Solving Answers
- 9. Balancing eBooks and Physical Books Holt Chemistry Stoichiometry Problem Solving Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Holt Chemistry Stoichiometry Problem Solving Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Holt Chemistry Stoichiometry Problem Solving Answers
  - Setting Reading Goals Holt Chemistry Stoichiometry Problem Solving Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Holt Chemistry Stoichiometry Problem Solving Answers
  - Fact-Checking eBook Content of Holt Chemistry Stoichiometry Problem Solving Answers
  - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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