

## UNIT III - THERMOCHEMISTRY

1. Calculate the H-F bond energy given the following data.



$$\Delta H_{\text{rxn}} = -544 \text{ kJ}$$

- A. 1130 kJ/mol    B. 565 kJ/mol    **C. 283 J/mol**    D. 300 kJ/mol    E. 134 kJ/mol
2. 1.0 g of the rocket fuel hydrazine ( $\text{N}_2\text{H}_4$ ) is burned in a bomb calorimeter. The temperature of the bomb calorimeter changes from 24.63 °C to 28.16 °C. The heat capacity of the bomb calorimeter is 837 J/K. Calculate the heat of combustion per mol of hydrazine in the bomb calorimeter.

- A. 95.0 kJ/mol**    B. 2.96 kJ/mol    C.  $3.55 \times 10^3$   
D. 92.4 J/mol    E.  $1.13 \times 10^5$  kJ/mol

3. Calculate the enthalpy change for the following reaction of nitric oxide NO with ozone
- $\text{O}_3$
- .



Given the following data:



- A. -1083.7 kJ    B. -693.8 kJ    C. -198.8 kJ    **D. -93.7 kJ**    E. -93.7 kJ
4. An endothermic reaction does no work on the surroundings. What values do you predict for the heat and energy change of this reaction?

- A.  $q < 0$ ,  $\Delta E > 0$     **B.  $q < 0$ ,  $\Delta E < 0$**     C.  $q > 0$ ,  $\Delta E > 0$   
D.  $q > 0$ ,  $\Delta E < 0$     E.  $q > 0$ ,  $\Delta E = 0$

5. M&M candies contain 209.0 Calories per 1.5 oz serving. The fat content accounts for 38.3% of these Calories. What is the energy of the fat content in kJ?

- A. 19.1 kJ    B. 874 kJ    **C.  $1.91 \times 10^{-2}$  kJ**  
D.  $3.35 \times 10^5$  kJ    E. 335 kJ

6. Calculate the heat released when 2.00 g of butane from a cigarette lighter is burned.



- A. -90.7 kJ    **B. -98.3 kJ**    C. -49.2 kJ    D. -108.4 kJ    E. -196.6 kJ
7. When Karl Kaverman adds chilled grog to a granite mug, 10.9 kJ of energy is transferred from the mug to the grog. If the mug has a mass of 625 g and an initial temperature of 25 °C, what is the final temperature of the mug? The specific heat of granite is 0.79 J/g·K.

- A. 2.9 °C    B. 14 °C    **C. 22 °C**    D. 25 °C    E. 47 °C

# General Chemistry Questions Answers

**Annelies Wilder-Smith**



## **General Chemistry Questions Answers:**

The book delves into General Chemistry Questions Answers. General Chemistry Questions Answers is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into General Chemistry Questions Answers, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to General Chemistry Questions Answers
- Chapter 2: Essential Elements of General Chemistry Questions Answers
- Chapter 3: General Chemistry Questions Answers in Everyday Life
- Chapter 4: General Chemistry Questions Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of General Chemistry Questions Answers. The first chapter will explore what General Chemistry Questions Answers is, why General Chemistry Questions Answers is vital, and how to effectively learn about General Chemistry Questions Answers.

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4. In chapter 3, this book will examine the practical applications of General Chemistry Questions Answers in daily life. The third chapter will showcase real-world examples of how General Chemistry Questions Answers can be effectively utilized in everyday scenarios.

5. In chapter 4, this book will scrutinize the relevance of General Chemistry Questions Answers in specific contexts. This chapter will explore how General Chemistry Questions Answers is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, this book will draw a conclusion about General Chemistry Questions Answers. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of General Chemistry Questions Answers.

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