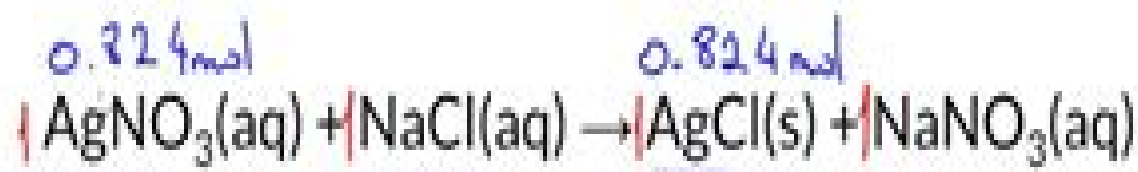


When 140.0 g of  $\text{AgNO}_3$  was added to an aqueous solution of  $\text{NaCl}$ , 86.0 g of  $\text{AgCl}$  was collected as a white precipitate. Which salt was the limiting reactant in this reaction? How many grams of  $\text{NaCl}$  was present in the solution?



1) L.R. - ?

2)  $m(\text{NaCl})$  - ?

$$n(\text{AgNO}_3) = \frac{m}{M} = \frac{140 \text{ g}}{169.9 \text{ g/mol}} = 0.824 \text{ mol}$$

$$n(\text{AgCl}) = \frac{m}{M} = \frac{86.0 \text{ g}}{143.3 \text{ g/mol}} = \underline{0.600 \text{ mol A.Y.}}$$

# Limiting Reactant Practice Problems With Answers

**John S. Phillips, Cheryl Wistrom**



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In a digitally-driven earth wherever screens reign supreme and instant connection drowns out the subtleties of language, the profound strategies and mental subtleties concealed within phrases frequently get unheard. However, located within the pages of **Limiting Reactant Practice Problems With Answers** a fascinating fictional value pulsing with raw feelings, lies an extraordinary journey waiting to be undertaken. Written by a skilled wordsmith, this marvelous opus invites visitors on an introspective trip, delicately unraveling the veiled truths and profound influence resonating within the fabric of every word. Within the emotional depths with this touching evaluation, we shall embark upon a genuine exploration of the book is primary styles, dissect its interesting writing fashion, and yield to the powerful resonance it evokes deep within the recesses of readers hearts.

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