Limiting Reactant

For the following reactions, find the following:

- a. Which of the reactants is the limiting reactant?
- b. What is the maximum amount of each product that can be formed?
- How much of the excess reactant is left over after the reaction is complete?
 SEL NEXT DOOK FOR COLCULATIONS!
- Consider the following reaction:

3 NH4NO3 + Na3PO4 → (NH4)3PO4 + 3 NaNO3

Answer the questions above, assuming we started with 30.0 grams of ammonium nitrate b and 50.0 grams of sodium phosphate.

30.09 NH4NO3 | 14001 NH4NO3 | 14001 (1414)3 PO4 | 149.129 (NH4)3 PO4 | 18.109 | 180.00 gNH4NO3 | 3 mol NH4NO3 | 14001 (149.129 (NH4)3 PO4 | 1414)3 PO4 | 1401 (149.129 (NH4)3 PO4 | 1414)3 PO4 | 1414 (149.129 (NH4)3 PO4 | 1414 (149.129 (

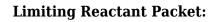
3 CaCO₃ + 2 FePO₄ → Ca₃(PO₄)₂ + Fe₂(CO₃)₃

Answer the questions at the top of this sheet, assuming we start with 100.0 graips of calcium carbonate and 45.0 grams of iron (II) phosphate.

100.0 g (a CO₃ | mol CaCO₅ | mol

Limiting Reactant Packet

Didier Musso



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