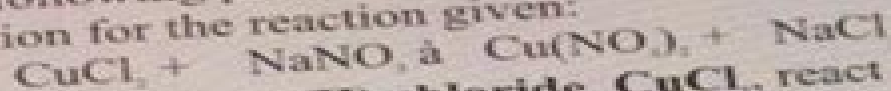


[ Limiting Reactant and Percent Yield Worksheet ]

Answer the following problems with complete solutions.

1. The equation for the reaction given:

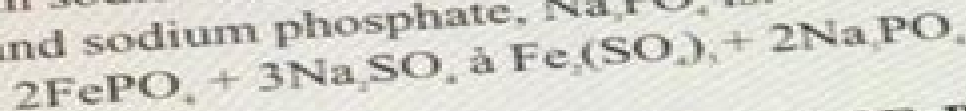


a) If 0.11 mol of copper (II) chloride,  $\text{CuCl}_2$ , react with 0.24 mol of sodium nitrate,  $\text{NaNO}_3$ , how many moles of sodium chloride,  $\text{NaCl}$ , can be formed?

b) What is the name of the limiting reagent?

c) If 11.3 grams of sodium chloride (actual yield) are formed in the reaction, what is the percent yield of this reaction?

2. The balanced equation for the reaction of iron (III) phosphate,  $\text{FePO}_4$ , with sodium sulfate,  $\text{Na}_2\text{SO}_4$ , to make iron (III) sulfate,  $\text{Fe}_2(\text{SO}_4)_3$ , and sodium phosphate,  $\text{Na}_3\text{PO}_4$ , is:



a) If you perform this reaction with 0.17 mol of  $\text{FePO}_4$  and an excess of  $\text{Na}_2\text{SO}_4$ , how many moles of  $\text{Fe}_2(\text{SO}_4)_3$  can you make?

b) If 18.5 grams of  $\text{Fe}_2(\text{SO}_4)_3$  are actually made when you do this reaction, what is your percent yield?

# Limiting Reactant And Percent Yield Lab Answers

**Rachel Sandford**



## **Limiting Reactant And Percent Yield Lab Answers:**

*Calculating Theoretical Yield Is Not Impossible!* Michael S. Leonard, 2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory. This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill. It begins with simple examples from everyday life, demonstrates the importance of balancing the equation, addresses the role of the mole in these computations, discusses different types of liquids, considers the role of significant figures, and culminates with the planning of syntheses. There are suggestions for further reading as well as practice problems and questions to ensure mastery. Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory. Addresses the necessity of balancing the reaction equation, the centrality of the mole in these calculations, and the role of significant figures in reporting the answer. Explains how to approach the calculations when using neat liquids or solutions. The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product. All of the problems in the book include detailed solutions with accompanying text to explain the answers, and ancillaries also include suggestions for further reading.

Chemistry Workbook For Dummies with Online Practice Chris Hren, Peter J. Mikulecky, 2017-03-21 Take the confusion out of chemistry with hundreds of practice problems. Chemistry Workbook For Dummies is your ultimate companion for introductory chemistry at the high school or college level. Packed with hundreds of practice problems, this workbook gives you the practice you need to internalize the essential concepts that form the foundations of chemistry. From matter and molecules to moles and measurements, these problems cover the full spectrum of topics you'll see in class, and each section includes key concept review and full explanations for every problem to quickly get you on the right track. This new third edition includes access to an online test bank where you'll find bonus chapter quizzes to help you test your understanding and pinpoint areas in need of review. Whether you're preparing for an exam or seeking a start to finish study aid, this workbook is your ticket to acing basic chemistry. Chemistry problems can look intimidating; it's a whole new language with different rules, new symbols, and complex concepts. The good news is that practice makes perfect, and this book provides plenty of it with easy-to-understand coaching every step of the way. Delve deep into the parts of the periodic table. Get comfortable with units, scientific notation, and chemical equations. Work with states, phases, energy, and charges. Master nomenclature, acids, bases, titrations, redox reactions, and more. Understanding introductory chemistry is critical for your success in all science classes, so following up with the material now makes life much easier down the education road. Chemistry Workbook For Dummies gives you the practice you need to succeed.

*Basic Concepts of Chemistry* Leo J. Malone, Theodore O. Dolter, 2011-12-27 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at

the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter WileyPLUS sold separately from text

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**A Concrete Stoichiometry Unit for High School Chemistry** Jennifer Louise Pakkala, 2006 **Forensics in Chemistry** Sara McCubbins, Angela Codron, 2012 Forensics seems to have the unique ability

to maintain student interest and promote content learning I still have students approach me from past years and ask about the forensics case and specific characters from the story I have never had a student come back to me and comment on that unit with the multiple choice test at the end from the Introduction to Forensics in Chemistry The Murder of Kirsten K How did Kirsten K s body wind up at the bottom of a lake and what do wedding cake ingredients soil samples radioactive decay bone age blood stains bullet matching and drug lab evidence reveal about whodunit These mysteries are at the core of this teacher resource book which meets the unique needs of high school chemistry classes in a highly memorable way The book makes forensic evidence the foundation of a series of eight hands on week long labs As you weave the labs throughout the year and students solve the case the narrative provides vivid lessons in why chemistry concepts are relevant and how they connect All chapters include case information specific to each performance assessment and highlight the related national standards and chemistry content Chapters provide Teacher guides to help you set up Student performance assessments A suspect file to introduce the characters and new information about their relationships to the case Samples of student work that has been previously assessed and that serves as an answer key for you Grading rubrics Using Forensics in Chemistry as your guide you will gain the confidence to use inquiry based strategies and performance based assessments with a complex chemistry curriculum Your students may gain an interest in chemistry that rivals their fascination with Bones and CSI

*Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter,1993*      **Chemical principles** Paul B. Kelter,1998      Chemistry James E. Brady,John R. Holum,1995-12-29 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping readers develop skills in analytical thinking and problems solving

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