

**Problem Solving** *continued*

- 12.** Urea,  $(\text{NH}_2)_2\text{CO}$ , is an important fertilizer that is manufactured by the following reaction.



What volume of  $\text{NH}_3$  at STP will be needed to produce  $8.50 \times 10^4$  kg of urea if there is an 89.5% yield in the process?

- 13.** An obsolete method of generating oxygen in the laboratory involves the decomposition of barium peroxide by the following equation.



What mass of  $\text{BaO}_2$  reacted if 265 mL of  $\text{O}_2$  is collected by water displacement at 0.975 atm and  $10.^\circ\text{C}$ ?

- 14.** It is possible to generate chlorine gas by dripping concentrated  $\text{HCl}$  solution onto solid potassium permanganate according to the following equation.



If excess  $\text{HCl}$  is dripped onto 15.0 g of  $\text{KMnO}_4$ , what volume of  $\text{Cl}_2$  will be produced? The  $\text{Cl}_2$  is measured at  $15^\circ\text{C}$  and 0.959 atm.

- 15.** Ammonia can be oxidized in the presence of a platinum catalyst according to the following equation.



The  $\text{NO}$  that is produced reacts almost immediately with additional oxygen according to the following equation.



If 35.0 kL of oxygen at STP react in the first reaction, what volume of  $\text{NH}_3$  at STP reacts with it? What volume of  $\text{NO}_2$  at STP will be formed in the second reaction, assuming there is excess oxygen that was not used up in the first reaction?

- 16.** Oxygen can be generated in the laboratory by heating potassium chlorate. The reaction is represented by the following equation.



What mass of  $\text{KClO}_3$  must be used in order to generate 5.00 L of  $\text{O}_2$ , measured at STP?

- 17.** One of the reactions in the Solvay process is used to make sodium hydrogen carbonate. It occurs when carbon dioxide and ammonia are passed through concentrated salt brine. The following equation represents the reaction.



- a.** What volume of  $\text{NH}_3$  at  $25^\circ\text{C}$  and 1.00 atm pressure will be required if 38 000 L of  $\text{CO}_2$ , measured under the same conditions, react to form  $\text{NaHCO}_3$ ?
- b.** What mass of  $\text{NaHCO}_3$  can be formed when the gases in (a) react with  $\text{NaCl}$ ?

# Holt Chemistry Stoichiometry Problem Solving Answers

**L Reisser**



## **Holt Chemistry Stoichiometry Problem Solving Answers:**

Holt Chemistry Ralph Thomas Myers, 2004      Problems and Problem Solving in Chemistry Education Georgios Tsaparlis, 2021 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner      **College Chemistry Questions and Answers PDF** Arshad Iqbal, The College Chemistry Quiz Questions and Answers PDF Class 11 12 Chemistry Competitive Exam Questions Chapter 1 6 Practice Tests Grade 11 12 Chemistry Textbook Questions for Beginners includes revision guide for problem solving with hundreds of solved questions Class 11 12 Chemistry Questions and Answers PDF book covers basic concepts analytical and practical assessment tests Class 11 12 Chemistry Quiz PDF book helps to practice test questions from exam prep notes The Grade 11 12 Chemistry Quiz Questions and Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved tests Class 11 12 Chemistry Questions and Answers PDF Free download chapter 1 a book covers solved common questions and answers on chapters atomic structure basic chemistry chemical bonding chemistry experimental techniques gases liquids and solids tests for college and university revision guide Chemistry Interview Questions and Answers PDF Download free eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The Class 11 12 Chemistry Interview Questions Chapter 1 6 PDF book includes college question papers to review practice tests for exams Class 11 12 Chemistry Practice Tests a textbook s revision guide with chapters tests for NEET MCAT GRE GMAT SAT ACT competitive exam College Chemistry Questions Bank Chapter 1 6 PDF book covers problem solving exam tests from chemistry textbook and practical eBook chapter wise as Chapter 1 Atomic Structure Questions Chapter 2 Basic Chemistry Questions Chapter 3 Chemical Bonding Questions Chapter 4 Experimental Techniques Questions Chapter 5 Gases Questions Chapter 6 Liquids and Solids Questions The Atomic Structure Quiz Questions PDF e Book Chapter 1 interview questions and answers on Atoms atomic spectrum atomic

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*Undergraduate Announcement* University of Michigan--Dearborn, 1989

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