

Worksheet #3—Stoichiometry

Directions: Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation: $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

$$\frac{14.8 \text{ moles Al}}{4 \text{ mole Al}} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole Al}} = \underline{11.1 \text{ mol O}_2}$$

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride (ZnCl_2) are produced?



$$\frac{20 \text{ g Zn}}{65 \text{ g Zn}} \times \frac{1 \text{ mol Zn}}{1 \text{ mole Zn}} \times \frac{1 \text{ mole ZnCl}_2}{1 \text{ mole Zn}} \times \frac{136 \text{ g ZnCl}_2}{1 \text{ mol ZnCl}_2}$$

$$= \underline{41.8 \text{ g ZnCl}_2}$$

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?



$$\frac{4.0 \text{ g AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mol AgNO}_3}{1 \text{ mole AgNO}_3} \times \frac{1 \text{ mole Cu}}{2 \text{ mole AgNO}_3} \times \frac{63.5 \text{ g Cu}}{1 \text{ mole Cu}}$$

2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO_3).



$$\frac{0.61 \text{ moles KClO}_3}{2 \text{ mole KClO}_3} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole KClO}_3} =$$

$$\underline{0.915 \text{ mol O}_2}$$

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed



$$\frac{10 \text{ g NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole Cl}_2}{2 \text{ mole NaCl}} \times \frac{71 \text{ g Cl}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{6.07 \text{ g Cl}_2}$$

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?



$$\frac{15.1 \text{ g Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mol Cl}_2}{1 \text{ mole Cl}_2} \times \frac{1 \text{ mole H}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{0.21 \text{ moles H}_2}$$

Holt Chemistry Problem Solving Work Answers

Stoichiometry

Thomas Griffiths



Holt Chemistry Problem Solving Work Answers Stoichiometry:

The Top Books of the Year Holt Chemistry Problem Solving Work Answers Stoichiometry The year 2023 has witnessed a noteworthy surge in literary brilliance, with numerous compelling novels captivating the hearts of readers worldwide. Let's delve into the realm of popular books, exploring the engaging narratives that have enthralled audiences this year.

The Must-Read : Colleen Hoover's "It Ends with Us" This touching tale of love, loss, and resilience has gripped readers with its raw and emotional exploration of domestic abuse. Hoover skillfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can succeed.

Holt Chemistry Problem Solving Work Answers Stoichiometry : Taylor Jenkins Reid's "The Seven Husbands of Evelyn Hugo" This spellbinding historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reid's compelling storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery.

Discover the Magic : Delia Owens' "Where the Crawdads Sing" This captivating coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens spins a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These bestselling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of captivating stories waiting to be discovered.

The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. *The Secret History* is an exceptional and gripping novel that will keep you guessing until the very end. The novel is a warning tale about the dangers of obsession and the power of evil.

<https://staging.conocer.cide.edu/results/Resources/Documents/key%20to%20physical%20science%20cpo%20answers.pdf>

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