

Activity A:
Alpha decay

Get the Gizmo ready:

- Click **Reset** (↺).
- Check that **Alpha decay** and **Uranium** are selected.



Question: How does alpha decay change the nucleus of a radioactive atom?

1. **Predict:** As you observed in the warm-up activity, an alpha particle consists of two protons and two neutrons. How will the emission of an alpha particle affect the following?

- A. The atomic number of the atom:
B. The mass number of the atom:

decrease by two

decrease by four

2. **Calculate:** Turn on **Write equation**. What you see is an equation that shows the original uranium atom on the left. The boxes on the right represent the **daughter product**—the atom produced by radioactive decay—and the emitted alpha particle.

- A. In the top left box, write the mass number of the daughter product and press "Enter" on your keyboard. What is this number?

234

- B. In the bottom left box, write the atomic number of the daughter product and press "Enter." What is this number?

90

- C. Click on the picture below, select edit. Enter the mass number and atomic number of the alpha particle, which has the same composition as the nucleus of a helium (He) atom. After filling in the boxes in the Gizmo, write the completed equation below:



- D. According to your equation, what **isotope** remains after the alpha decay of uranium-238? (Note: You can look up element symbols on the periodic table.)

Th - 234

3. **Check:** Turn on **Show equation**, and click **Play**. The equation will appear at the end of the animation. Was your prediction correct? **Yes** If not, modify your equation above.

Key Terms Nuclear Changes Answer Key

**Theodore L. Brown, H. Eugene LeMay
Jr., Bruce E. Bursten, Catherine
Murphy, Patrick Woodward, Steven
Langford, Dalius Sagatys, Adrian
George**

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