

Min. X/R Ratios to Test Circuit Breakers

$$PF = \cos [\tan^{-1}(X/R)]$$

Protective Devices	Type of equipment	First cycle current	First cycle X/R	Interrupting time or short time X/R
	Low-voltage power circuit breaker (iron frame breaker)	Peak	6.59	6.59
	Low-voltage molded and insulated case breakers with interrupting ratings > 20 kA	Peak	4.9	4.9
	Low-voltage molded and insulated case breakers with interrupting ratings 10-20 kA	Peak	3.18	3.18
	Low-voltage molded and insulated case breakers with interrupting ratings < 10 kA	Peak	1.73	1.73
	Fused low-voltage power circuit breakers	Peak	4.9	4.9
	Low-voltage busway	Peak	4.9	4.9
	High-voltage power circuit breaker	rms	25	15 ¹⁷
	Power fuse	rms	15	—
	Distribution fuse	rms	10	—
	Distribution air cutout fuse	rms	5-15 depends on kV rating	—
	Distribution oil cutout fuse	rms	9-12 depends on kV rating	—
	Switches (withstand rating)	rms	25	—
	High-voltage bus duct		25	—

Iec Standard For Circuit Breaker

**Hugh McLaren Ryan, Institution of
Electrical Engineers**



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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm³ 3 place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol⁻¹

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic eg 1 HCl(aq) + NaOH(aq) → NaCl(aq) + H₂O(l) for which ΔH = 57 kJ mol⁻¹ eg 2 HNO₃(aq) + KOH(aq) → KNO₃(aq) + H₂O(l) for which ΔH = 57 kJ mol⁻¹

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example HCl(aq) + NaOH(aq) → NaCl(aq) + H₂O(aq) you can see that one mole of water is produced when one mole of HCl or one mole of NaOH reacts