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# **Holt Geometry Teaching Transparency**

**Emilie Sanchez** 

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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1s orbital and the third is placed in the 2s orbital giving an electron configuration of 1s 2s 2s 1

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group iia

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web the electron removed during the ionization of beryllium he 2s 2 is an s electron whereas the electron removed during the

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is 1s 2 2s 2 2p 6 3s 2 3p 6 4s 1 or simply put ar 4s 1 q2 what is the total number of electrons that can be filled in s p and d subshells

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web by building up from hydrogen this table can be used to determine the electron configuration for any atom on the periodic table we will now construct the ground state electron configuration and orbital diagram for a selection of atoms in the first and second periods of the periodic table

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