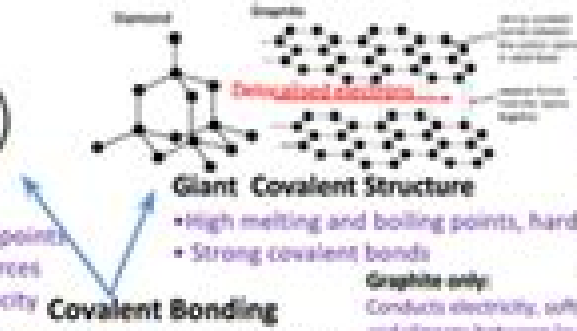
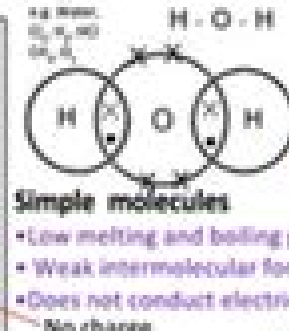




Periodic Table of the Elements

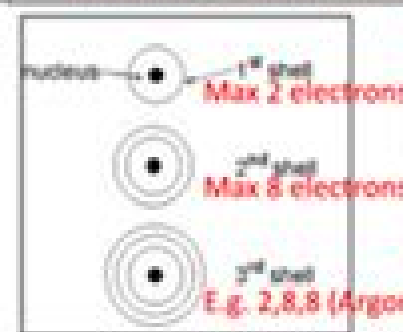
- metals
- non-metals
- transition metals
- alkali metals
- alkali earth metals
- transition metals
- lanthanoids
- actinoids
- halogens
- noble gases
- metalloids

An atom:  
Number of Protons = Number of Electrons  
Mass number = Protons + Neutrons

# C2 Chemistry

Periodic table arranges elements by  
**ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.  
E.g. Lithium 2,1 is a group 1 element.

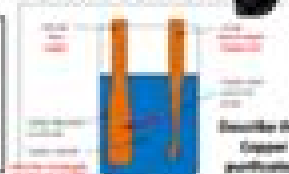


Protons and neutrons both have a mass of 1

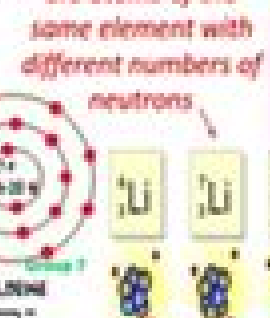
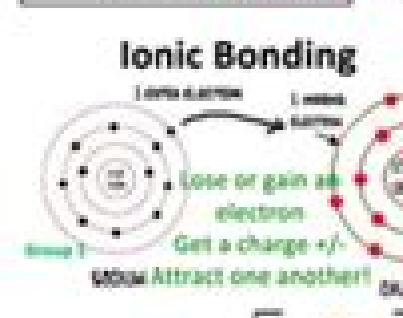
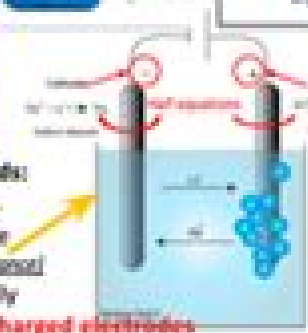
**ISOTOPES**  
are atoms of the same element with different numbers of neutrons

**OIL RIG**

Oxidation is loss (of electrons)  
Reduction is gain (of electrons)



**Electrolysis:**  
Splitting up a substance using electricity



When there are 2 possible products  
The less reactive forms at the electrode

Electrolysis of Brine

Used for: Disinfectants, Bleach and plastics

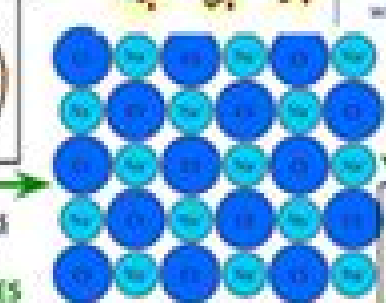
Used for: Margarine and HCl

Used for: Bleach, paper and soap

Metal ions only arranged in layers  
**ELECTROSTATIC ATTRACTION**  
between  
**+IONS and ELECTRONS**

Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well



As a solid: High melting and Boiling Point  
Conducts electricity when melted or dissolved:  
Ions move freely and carry a charge

**NANO-TECHNOLOGY**  
Means **REALLY REALLY REALLY** small!!!!

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employment status which was terminated due to my illness injury i have fully recovered and am excited to return to work i  
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