

A chemical bond - forms when 2 or more atoms rearrange valence
electrons to increase stability.

ionic bond - forms when valence electrons are transferred
from one atom to another

cation - atom loses electrons to become + charged

anion - atom gains electrons to become - charged

In ionic compounds the ions are arranged in a crystal lattice.
Strong forces hold the ions together.

Properties of ionic compounds:

- high melting and boiling points
- hard - not easily crushed
- conduct electricity when melted or
dissolved because the ions are free to move.

covalent bond - electrons are shared, forming
molecules.

Covalent compounds have weaker forces holding the
molecules together.

Properties of covalent compounds:

- lower melting and boiling points
- Many covalent compounds are volatile liquids or gases.
- soft - easier to crush
- are not conductors of electricity

electronegativity - property that tells how strong an atom's attraction
is for electrons.

Since oxygen has a higher electronegativity than hydrogen, oxygen
holds onto shared electrons tighter, giving the oxygen a partial
negative charge and the hydrogen a partial positive charge.

Ionic Bonds And Notetaking Guide Answer

B Lingard



Ionic Bonds And Notetaking Guide Answer:

Decoding **Ionic Bonds And Notetaking Guide Answer**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Ionic Bonds And Notetaking Guide Answer**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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