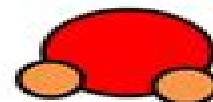


Chapter 6: Chemical Bonding Guided Notes

I. Why do Atoms Combine?

A. Chemical Formula

1. The chemical formula shows:
 - a. The _____ in the compound.
 - b. The ratio of their _____

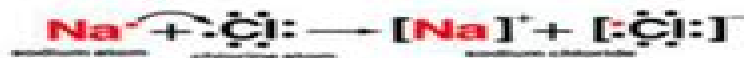


B. Chemical Bond

1. Strong attractive force between _____ or ions in a molecule or compound.
2. Formed by:
 - a. _____ electrons, (e-) by losing or gaining electrons.
 - b. _____ electrons (e-)

C. Stability

1. Octet Rule
 - a. Most atoms form bonds in order to have _____ valence electrons (e-).
 - b. _____ outer energy level.
 - c. Want to be like the Noble Gases (group 18).
2. Stability is the driving force behind bond formation. All atoms bond to become stable.
3. Transferring e-



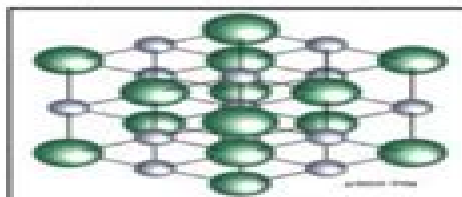
4. Sharing e-



II. Kinds of Chemical Bonds

A. Ionic Bond

1. Attraction between 2 _____ charged ions.
2. Ions are _____ atoms.
3. A _____ is a _____ charged ion. Forms when an atom loses an electron.
Formed by _____.
4. A _____ is a _____ charged ion. Forms when an atom gains an electron.
Formed by _____.
5. Formed by _____ electrons (e-) from a _____ to a _____.
6. Ions form a 3-D crystal lattice.



Guided Notes Ionic Bonds Chapter 6

Arshad Iqbal



Guided Notes Ionic Bonds Chapter 6:

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In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Guided Notes Ionic Bonds Chapter 6**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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