

KEY

Ionic Formula Practice

Name	Cation	Anion	Formula
1. calcium bromide	Ca^{+2}	Br^{-1}	Ca Br_2
2. magnesium phosphide	Mg^{+2}	P^{-3}	$\text{Mg}_3 \text{P}_2$
3. aluminum oxide	Al^{+3}	O^{-2}	$\text{Al}_2 \text{O}_3$
4. barium sulfide	Ba^{+2}	S^{-2}	Ba S
5. sodium phosphide	Na^{+1}	P^{-3}	$\text{Na}_3 \text{P}$
6. strontium chloride	Sr^{+2}	Cl^{-1}	Sr Cl_2
7. calcium nitride	Ca^{+2}	N^{-3}	$\text{Ca}_3 \text{N}_2$
8. ammonium bromide	NH_4^{+1}	Br^{-1}	$\text{NH}_4 \text{Br}$
9. calcium sulfate	Ca^{+2}	SO_4^{-2}	Ca SO_4
10. magnesium nitrate	Mg^{+2}	NO_3^{-1}	$\text{Mg}(\text{NO}_3)_2$
11. sodium carbonate	Na^{+1}	CO_3^{-2}	$\text{Na}_2 \text{CO}_3$
12. zinc hydroxide	Zn^{+2}	OH^{-1}	$\text{Zn}(\text{OH})_2$
13. barium nitrite	Ba^{+2}	NO_2^{-1}	$\text{Ba}(\text{NO}_2)_2$
14. aluminum sulfite	Al^{+3}	SO_3^{-2}	$\text{Al}_2(\text{SO}_3)_3$
15. lead II nitrate	Pb^{+2}	NO_3^{-1}	$\text{Pb}(\text{NO}_3)_2$
16. copper I phosphate	Cu^{+1}	PO_4^{-3}	$\text{Cu}_3 \text{PO}_4$
17. ammonium carbonate	NH_4^{+1}	CO_3^{-2}	$(\text{NH}_4)_2 \text{CO}_3$
18. tin IV nitrate	Sn^{+4}	NO_3^{-1}	$\text{Sn}(\text{NO}_3)_4$
19. nickel II acetate	Ni^{+2}	$\text{C}_2\text{H}_3\text{O}_2^{-1}$	$\text{Ni}(\text{C}_2\text{H}_3\text{O}_2)_2$
20. chromium III sulfate	Cr^{+3}	SO_4^{-2}	$\text{Cr}_2(\text{SO}_4)_3$
21. iron III nitrite	Fe^{+3}	NO_2^{-1}	$\text{Fe}(\text{NO}_2)_3$
22. cobalt II hydroxide	Co^{+2}	OH^{-1}	$\text{Co}(\text{OH})_2$

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Mike Jess

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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