

Student Answer Keys (also included page 2) The Online Learning Center helps you learn English, Essential Questions Edition. You can find course-wide content, such as classroom questions and an interview process, on the left navigation bar. You can also explore a chapter from the course to access chapter-specific content, such as questions and writing exercises. The Online Learning Center provides you with secure, downloadable resources. You can find course-wide content, such as PowerPoint presentations and instructor's manuals, on the left navigation bar. **Notes**

Chapter 8 Atomic Structure Section 8.1 Studying Atomic (pages 189-195) This section explains how atomic models evolved. **Reading Strategy (page 190)** Summarizing As you read, fill in the table about atomic models. For more details on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook. **Atomic Models Scientific Evidence Model** Model of atoms in compounds Deducted from Rutherford's Positive, dense nucleus. **Atomic Greek Models of Atoms (page 191)** 1. Democritus called the smallest particles of matter because they could not be split. 2. He saw the four elements that Aristotle used in his model of matter: a, b, c, d. Dalton's Atomic Theory (page 193) 3. Thomson's Model

Model	Structure
1. Democritus's Model	Democritus proposed that matter is made of very small particles called atoms. Atoms cannot be created or destroyed. Atoms are indivisible particles, which cannot be created or destroyed.
2. Dalton's Atomic Theory	Dalton's Atomic Theory is based on the following postulates: 1. Matter is made of very small particles called atoms. 2. Atoms cannot be created or destroyed. 3. The relative number and kinds of atoms are constant in a chemical reaction. 4. The ratio of the number of atoms of each element in a reaction is constant.
3. Thomson's Model	Thomson's model of an atom is a sphere of positive charge with electrons embedded in it. The sphere is called a plum pudding model.
4. Rutherford's Model	Rutherford's model of an atom is a central nucleus with a positive charge. Electrons are present outside the nucleus in different shells or orbits.
5. Bohr's Model	Bohr's model of an atom is a central nucleus with a positive charge. Electrons are present outside the nucleus in different shells or orbits. The electrons move in fixed orbits.

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