



Electronic Configuration



1. Use noble gases to write the abbreviated electron configurations of the following?

i. Li _____

ii. B _____

iii. O _____

iv. Mg _____

v. Cl _____

vi. V _____

vii. Se^{2-} _____

viii. Cr^{3+} _____

ix. I _____

x. Ag _____

2. Identify the element with the following electron configuration

i. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$ _____

ii. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$ _____

iii. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^7$ _____

iv. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10}$ _____

v. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^1$ _____

3. Determine if the following electron configurations are correct:

i. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^6 5s^1$ _____

ii. $1s^2 2s^2 2p^6 3s^3$ _____

iii. $[\text{Rn}] 7s^2 5f^8 6d^2$ _____

iv. $[\text{Ar}] 5s^2 4d^{10} 5p^5$ _____

v. $[\text{Xe}] 6s^2 4f^{10}$ _____

4. Identify the following elements

i. $1s^2 2s^2 2p^2$ _____

ii. $1s^2 2s^2 2p^6$ _____

iii. $[\text{Ar}] 4s^2 3d^{10} 4p^5$ _____

iv. $[\text{Kr}] 5s^2 4d^1$ _____

v. $[\text{Ne}] 3s^2 3p^3$ _____

Electron Configuration Worksheet Pogil Activities For High School Chemistry

Mark Steyvers



Electron Configuration Worksheet Pogil Activities For High School Chemistry:

Electron Structure and High-temperature Chemistry of Coordination Compounds, 1996 *Atomic and Molecular Electronic Configuration Revisited* Harold J. Teague, 2014-06-11 In AMEC chapter one you will find a structure for O₂ supported by both bond order and unpaired electron data Previous attempts at providing an electronic structure for O₂ fails due to bond order number of unpaired electrons or oxygen atoms indicating five orbitals although four is the limit Herein you will find the newly described LONE P ORBITAL Pi BOND which is the solution not only for O₂ but also for NO and OF uncharacterized The LPOP bond system is possible due to the small size and high electronegative character of the N O and F atoms which makes it related to H Bonding Diatomic structures involving row two periodic table elements LiBe Be₂ BeB B₂ BC C₂ CN N₂ and F₂ are also presented All structures are shown being formed by two methods MRAE Most Recently Added Electron and LCAO Linear Combination of Atomic Orbitals Chapter one also includes a Section III MRAE Mechanistic Sequence For NO NO O₂ OF F₂ and F₂ F F₀ Section IV Photographs of NO and O₂ molecular models Section V MRAE MO Energy Level diagrams for the diatomic structures listed above and Section VI a MRAE MO Energy Level Diagrams Template which is based on the cascade of changes in electronic distribution when an electron is forced onto a previous structure MRAE such as N₂ going to N₂ and NO going to NO Chapter 2 includes sections on hybridization and mechanism as it relates to polyatomic species The following sequences are given C H CH H CH₂ H CH₃ H CH₄ saturated and tetrahedral arrangement and N H NH H NH₂ H NH₃ H NH₄ also saturated and tetrahedral arrangement The emphasis is to show trend from atomic structure through sp sp² and sp³ hybridization Chapter 3 is composed of selections from Atomic and Molecular Configuration 1987 Included is MRAE AUFBAU TEMPLATE of the periodic table elements expanded to two pages It is informative to compare the MRAE concept for atoms to the MRAE method for diatomic species in Chapter one In both cases the key is the changes in electronic structure when an electron is added

ATOMIC STRUCTURE High School Chemistry Curriculum Bud Morgan, 2025-07-09 Welcome to our comprehensive guide to Atomic Structure designed specifically for high school students studying chemistry In this book we will explore the fascinating world of atoms the basic building blocks of matter and delve into the complexities of their structure Chapter 1 provides a solid foundation by introducing the concept of atoms and their historical significance in the development of modern chemistry We will examine the key scientists who contributed to our understanding of atomic structure from Democritus to Thomson Rutherford and Bohr Moving on to Chapter 2 we will explore the structure of the atom in more detail including the subatomic particles that make up an atom protons neutrons and electrons We will also discuss the concept of atomic number atomic mass and isotopes In Chapter 3 we will delve into the arrangement of electrons within an atom s electron cloud including energy levels sublevels and orbital shapes We will also discuss the principles of electron configuration and how it impacts an atom s chemical properties Chapter 4 will focus on the periodic table and how it organizes elements based on their atomic structure

We will explore periodic trends such as atomic radius ionization energy and electron affinity and discuss how these trends are related to an element's position on the table. In our final chapter we will discuss the significance of atomic structure in various chemical reactions including bonding stoichiometry and the behavior of gases. We will also explore the importance of understanding atomic structure in fields such as biochemistry environmental science and materials science. Throughout this book we will provide clear explanations helpful illustrations and interactive activities to reinforce key concepts. Whether you are a visual learner a hands on learner or a theoretical learner this book is designed to meet your needs. We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry. So let's dive in and unlock the mysteries of the atom together.

Introduction to the Electron Theory of Small Molecules Andrew Crowther Hurley, 1976

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