1.	For each	of the follow	ving sets of	sublevels,	determine	which has	the lowest
	energy, as	nd which ha	s the highe	est energy.			

(a) 2s, 3s, 4s (b) 4f, 4p, 4d (c) 4d, 4f, 3d (d) 6s, 3d, 5f Ex for (a) Lowest 2s, highest 4s

Explain what 6s<sup>2</sup> means.

3. What is the maximum number of electrons that can be found in an s - sublevel? p - sublevel? d - sublevel? f - sublevel?

4. Indicate which of these (there may be more than one) orbital designations is incorrect:

(a) 4s

(b) 3f

(c) 2d

(d) 3d

Draw energy level diagrams for

(a) sodium

(b) silicon

(c) beryllium

(d) cobalt

6. How many unpaired electrons are there in each of the atoms in question 5?

Give the electron configuration for:

(a) lithium

(b) oxygen (c) selenium

(d) iodine

(e) silver (f) arsenic (g) indium

(h) potassium

Give the name and symbol for the atom whose electron configuration corresponds to each of the following electron configurations.

a) 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>5</sup>

b) 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup> 4s<sup>2</sup>3d<sup>10</sup>4p<sup>3</sup>

c) 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup> 4s<sup>2</sup>3d<sup>10</sup>4p<sup>6</sup>5s<sup>2</sup>4d<sup>6</sup>

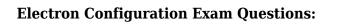
d) 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>2</sup>

e) 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup> 4s<sup>2</sup>3d<sup>10</sup>4p<sup>6</sup>5s<sup>1</sup>

9. Write the shorthand notation for the above electron configurations.

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