

Acid Base Homework

Exercise 16.69

Part A

If 20.0 mL of glacial acetic acid (pure $\text{HC}_2\text{H}_3\text{O}_2$) is diluted to 1.50 L with water, what is the pH of the resulting solution? The density of glacial acetic acid is 1.05 g/mL.

Express your answer to two decimal places.



ATP



$\text{pH} =$ 3.63

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Lingsheng Yao



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