Chapter 19 Practice Problem Set

- 1. Determine whether each of the following process is spontaneous or non-spontaneous.
 - a. Spontaneous / nonspontaneous dissolving table salt in hot soup
 - b. Spontaneous / nonspontaneous ball rolling up A Mountain
 - Spontaneous / nonspontaneous opening a perfume bottle to allow the fragrance to spread through the room
 - d. Spontaneous / nonspontaneous separating helium and neon from a mixture of gases
 - e. Spontaneous / nonspontaneous water solidifying to ice at 3 °C
- Which one of the following processes produces a decrease in the entropy of the system?
 - A. boiling water to form steam
 - B. dissolution of solid KCI in water
 - C. mixing of two gases into one container
 - D. freezing water to form ice
 - E. melting ice to form water
- Which of these species has the largest entropy (5") at 25°C?
 - A) CO(g) B) CHa[g] C) NaCl(s) D) H₂O(l) E) Fe(s)
- 4. For each of the following pairs, identify which one has a greater number of microstates (W):
 - a. CH₄ or C₂H₄
 - b. H₂O₂₈ or H₂O₃₀
 - c. 1 mol O₂ occupying 1L a 25°C or 1 mol O₂ occupying 2L at 25°C
- Predict whether each of the following reactions would have a positive or negative entropy change:
 - a. Δ5 ≥ 0 S₍₀₎ + O₂₍₀₎ → SO₂₍₀₎
 - b. $\Delta S \leq 0$ $H_2O_{BS} \rightarrow H_2O_{BS}$
 - c. Δ5 < 0 PCl_{3 M} + Cl_{2 M} → PCl_{3 M}
 - d. $\Delta S > 0$ $2Na_{(0)} + 2H_2O_{(1)} \rightarrow 2NaOH_{(ac)} + H_{2(0)}$
 - e. AS ≥ 0 N_{2(g)} → 2N_g

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