

Page 5	Mark Scheme IGCSE – May/June 2013	Syllabus 0625	Paper 31
3	<p>(a) (i) 1. (loss of P.E. =) mgh OR $92 \times 10 \times 1500$ $1.38 \times 10^6 \text{ J}$ correct use of mgh with $h = 500$ or 2000 gains 1 mark only</p> <p>(ii) 2. (K.E. =) $\frac{1}{2}mv^2$ OR $\frac{1}{2} \times 92 \times 52^2$ $1.244 \times 10^5 \text{ J}$ at least 2 sig. figs</p> <p>(a) (ii) difference is due to: (work done in overcoming) air resistance/drag OR energy converted to/lost as heat (by air resistance/drag)</p> <p>(b) (i) increases</p> <p>(ii) 920 N</p>		<p>C1 A1</p> <p>C1 A1</p> <p>B1</p> <p>B1</p> <p>B1</p> <p>[Total 7]</p>
4	<p>(a) (i) mention of vacuum OR glass is a poor conductor OR vacuum/gap between walls has no molecules/atoms/particles</p> <p>(ii) surface/silver (of walls) is good reflector/poor absorber (of radiation) surface/silver (of walls) is poor emitter (of radiation)</p> <p>(b) add a stopper/lid/bung/cover/top to reduce/prevent (loss of heat by) convection/ conduction/radiation/evaporation OR to prevent steam/hot vapour leaving</p> <p>made of insulator OR example of insulator to reduce/prevent (loss of heat by) convection/radiation/evaporation OR to prevent steam/hot air leaving</p>		<p>B1</p> <p>B1 B1</p> <p>M1 B1</p> <p>B1</p> <p>[Total 6]</p>
5	<p>(a) (i) and (ii) marked together to maximum of 3 marks</p> <p>(i) molecules escape/leave the liquid/form gas or vapour</p> <p>(ii) evaporation OR heat/(thermal) energy needed for evaporation leaves sweat cooler fast(er) molecules/high(er) energy molecules escape OR slow(er) molecules left behind heat flows from body to warm the sweat (so body cools)</p> <p>(b) (i) ($Q =$) $mc\Delta\theta$ OR mcT OR $60 \times 4000 \times 0.50$ $1.2 \times 10^5 \text{ J}$ / 120 kJ</p> <p>(ii) $Q = mL$ in any form OR ($m =$) Q/L OR either with numbers ($m = 1.2 \times 10^5 / 2.4 \times 10^5 =$) 0.05 kg e.c.f from (b)(i)</p>		<p>B1</p> <p>B1 B1 B1</p> <p>C1 A1</p> <p>C1 A1</p> <p>[Total 7]</p>

Mark Scheme 0625 61 0 N 1

Minjie Lin

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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web the arrhenius theory of acids and bases acids are substances that produce hydrogen ions in solution bases are substances that produce hydroxide ions in solution neutralization happens because hydrogen ions and hydroxide ions react to produce water $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ hydrochloric

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web aug 15 2020 the equilibrium formulated in scheme 1 between hydrogen ion and the corresponding acid and base can be called a simple acid base equilibrium by mixing two simple systems a double acid base system and an acid base equilibrium result that can always be formulated as follows $\text{acid}_1 + \text{base}_2 \rightleftharpoons \text{acid}_2 + \text{base}_1$

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web the A^- ion is a base because it is accepting a proton from the H_2O the reversible reaction contains two acids and two bases we think of them in pairs called conjugate pairs when the acid HA loses a proton it forms a base A^- when the base A^-

accepts a proton back again it obviously reforms the acid

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