

1. According to the Brønsted—Lowry definition, which chemical species can function *both* as an acid and as a base?

- Cl^-
- SO_4^{2-}
- NH_4^+
- HCO_3^-
- H_3O^+

2. In the reaction $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$ which is an acid—base conjugate pair?

- H_2O and HCN
- H_2O and OH^-
- CN^- and H_2O
- HCN and OH^-

3. Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX

- It is less basic than a 1 M solution of NaY.
- It is more basic than a 1 M solution of NaY.
- It yields a neutral solution.
- It is more concentrated than a 1 M solution of NaY.

4. The pH of a 0.03 M HCl solution is

a.	1.5
b.	2.5
c.	3.5
d.	12.5

5. The pH of a 1.0×10^{-3} M $\text{Ba}(\text{OH})_2$ solution at 25°C is

a.	2.7
b.	3.0
c.	11.0
d.	11.3

4. Which is the strongest acid?

- HClO
- HClO_2
- HClO_3
- HClO_4

7. Which salt reacts with water (hydrolyzes) to produce a basic solution?

a. $\text{NaC}_2\text{H}_3\text{O}_2$
b. NaNO_3
c. NH_4Cl
d. BaSO_4

8. The weakest of the bases listed is

Acid	Conjugate Base	K _a (Ionization Constant of Acid)
HCl	Cl ⁻	100% ionized
HSO ₄ ⁻	SO ₄ ²⁻	1.2 x 10 ⁻²
H ₂ S	HS ⁻	5.7 x 10 ⁻⁸
HS ⁻	S ²⁻	1.2 x 10 ⁻¹⁴

a.	Cl^-
b.	CN^-
c.	HS^-
d.	S^{2-}
e.	SO_4^{2-}

Year	Atlantic croaker	Atlantic menhaden	Atlantic herring	Atlantic bluefish
1960	45	15	10	5
1965	40	20	15	10
1970	35	25	20	15
1975	30	30	25	20
1980	25	35	30	25
1985	20	40	35	30
1990	15	45	35	25

10. Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate $\text{KC}_2\text{H}_3\text{O}_2$ is less alkaline than a 0.10 M solution of potassium cyanide, KCN?

- Hydrocyanic acid is a weaker acid than acetic acid.
- Hydrocyanic acid is less soluble in water than acetic acid.
- Cyanides are less soluble than acetates.
- Acetic acid is a weaker acid than hydrocyanic acid.

11. In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of

- 0.100 M sodium solution,
- 0.0500 M sodium hydroxide solution,
- 0.0500 M benzoic acid solution,
- 0.0500 M sodium benzoate solution.

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David Baud



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