

Equation Writing and Predicting Products

Introduction:

If you examine a car that has been in a junkyard for a while, you will notice that it has rusted. Rusting is a slow chemical reaction of the iron in the car's body with oxygen gas. If a piece of sodium is put into water, a much more rapid chemical reaction occurs. Sodium reacts with water to produce sodium hydroxide and hydrogen gas. During this reaction, enough heat is liberated to ignite the hydrogen gas, causing it to explode.

Chemists observe what is happening in a chemical reaction and try to describe it in language that is simple and clear. A chemical equation uses formulas and symbols to describe the substances involved in a reaction, the physical state of the substance, the use of a catalyst, and relative proportions. The general form of an equation is:



In this investigation you will perform a series of reactions and make careful observations of the changes that occur. Using simple tests and your knowledge of chemistry, you will determine the identity of the products. With this information, you will write chemical equations to describe the reactions.

Pre-Lab Discussion:

Read through the entire lab investigation and the relevant pages of your textbook. Then answer the questions that follow.

1. What constitutes a positive test for each of the following gases?
 - a. oxygen _____
 - b. hydrogen _____
 - c. carbon dioxide _____
 - d. water vapor _____
 - e. ammonia (NH_3) _____
2. What is the proper way to smell a substance in the lab? Why should care be taken when smelling a gas such as ammonia? _____
3. What is the role of a catalyst in a reaction? How can you tell when a substance serves as a catalyst? (This answer will be important later, spend some time on it now.) _____
4. One way to identify limestone (CaCO_3) is to drop a small amount of hydrochloric acid (HCl) on it. A positive test results in a fire of carbon dioxide being produced. Write out in words the information represented by the following balanced chemical equation.



Equation Writing And Predicting Products Lab Packet

Michael Hoelscher



Equation Writing And Predicting Products Lab Packet:

Laboratory Manual of General Chemistry Arthur Becket Lamb, 1916 Prentice Hall Chemistry, 2000

Laboratory Manual to Accompany Introductory Chemistry Uno Kask, Otto W. Nitz, Martin G. Ondrus, Tammy J. Melton, 1993-07 This fifth edition of this laboratory manual emphasizes safety in the lab and discusses equipment requirements in the apparatus section at the beginning of each experiment It also features a revised art programme and explains the rationale for each experiment Lab Manual for General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, 2009-08-21 Teaching all of the necessary concepts within the constraints of a one term chemistry course can be challenging Authors Denise Guinn and Rebecca Brewer have drawn on their 14 years of experience with the one term course to write a textbook that incorporates biochemistry and organic chemistry throughout each chapter emphasizes cases related to allied health and provides students with the practical quantitative skills they will need in their professional lives Essentials of General Organic and Biochemistry captures student interest from day one with a focus on attention getting applications relevant to health care professionals and as much pertinent chemistry as is reasonably possible in a one term course Students value their experience with chemistry getting a true sense of just how relevant it is to their chosen profession To browse a sample chapter view sample ChemCasts and more visit www.whfreeman.com/gob Laboratory Manual Arranged to Accompany "Principles of General Chemistry", by Stuart R. Brinkley and Erwin B. Kelsey Stuart Robert Brinkley, Erwin Burr Kelsey, 1926 A Laboratory Manual of General Chemistry for Use in Colleges William Crowell Bray, Wendell Mitchell Latimer, 1921 *ChemDiscovery Teacher Edition* Olga I. Agapova, 2002 Laboratory Manual of Physical Pharmaceutics I Dr. Vishwajeet S. Ghorpade, Dr. Remeth J. Dias, Dr. Kailas Mali, Dr. Prasad Patrekar, 2023-10-07 We are very pleased to put forth the first edition of Laboratory Manual of Physical Pharmaceutics I This manual is prepared as per PCI Education Regulations 2014 for Degree Course in Pharmacy This manual is designed for outcome based education and each experiment is arranged in a uniform way such as practical significance practical outcomes PrOs and its mapping with course outcomes minimum theoretical background resources used procedure precautions observations result conclusion references and related questions A sincere attempt has been made through this manual to provide practical knowledge to the students related to various experiments in Physical Pharmaceutics I The manual mainly includes the experiments through which the students will learn how the basic properties of drug like solubility dissolution can limit the bioavailability of drug and in turn the efficacy of the drug The solubility enhancement techniques like reduction of particle size use of solubilizing agents or prodrugs etc are also discussed The students will also learn about the importance of drug stability and how the parameters like temperature pH humidity and light can affect the drug stability Physical pharmaceutics is very important subject which helps the researchers to design and formulate various dosage forms including its manufacturing aspects The students will be proficient in handling various equipment used in physical pharmaceutics laboratory The students will also be able to

determine pKa partition coefficient surface tension critical micelle concentration HLB value stability constant buffer capacity etc in the laboratory We have taken extra efforts to make the experiments simpler easy to understand and perform The graphs have also been provided wherever necessary Each experiment is divided into sections like aim practical significance relevant course outcomes practical skills relevant affective domain related outcomes practical outcomes minimum theoretical background requirements related questions and references The manual has been designed with more emphasis on the practical skill improvement of the students so that the students can perform the practical with ease and comfort We are very much thankful to the designer publisher printers and all the stakeholders for putting their efforts for successfully bringing this manual out for the students Hope this manual will help the students to learn the concept principles and perform the experiments We wish them all the best *Laboratory Manual [for] Fundamentals of Chemistry* Fred H. Redmore,1979

The Success Manual for General Chemistry Elizabeth Kean,Catherine Middlecamp,1986 **Laboratory Manual of Organic Medicinal Products** Walter Henry Hartung,Wooten Taylor Sumerford,Melvin F. W. Dunker,1946 Index & Curriculum Briefs Curriculum materials clearinghouse,1974 Laboratory Manual of Chemistry, for Secondary Schools William Albine Morse,Frederick Charles Irwin,1905 **Laboratory Manual to Accompany Principles of General Chemistry** Erwin Burr Kelsey,Harold Glenn Dietrich,1933 **Chemistry**, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry The 100 Series science books span grades 5 to 12 The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics The series will be aligned to current science standards *Laboratory Manual* Jo A. Beran,1990 *Introduction to Chemistry, Laboratory Manual* T. R. Dickson,1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities These themes involve and engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students **A Laboratory Manual of General Chemistry** John Henry Secrist,1966 Instructor's Resource Manual Silberberg,1999-07

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