

1. Write the chemical equation for the reversible reaction studied in this experiment:



2. Write the equilibrium expression for the reaction studied in this experiment:

$$K_c = \frac{\text{FeSCN}^{2+}(\text{aq})}{(\text{Fe}^{3+})_{\text{aq}}(\text{SCN}^{-})_{\text{aq}}}$$

3. Complete the following table from your observations of test tubes #1-#6:

Test Tube	Change	Initial Color	Final Color (darker/lighter)
	Initial color before anything is added		
1	Added $\text{Fe}(\text{NO}_3)_3$	Reddish color	darker
2	Added KSCN	Reddish color	darker
3	Added NH_4Cl	Reddish color	lighter
4	Added AgNO_3	Reddish color	lighter
5	Cold	Reddish color	darker
6	Hot	Reddish color	lighter

4. Complete the following table for test tubes #2 - 4:

Test Tube	What concentration was changed and how it was changed?	The equilibrium shifted to left/right/no change	What evidence for the direction of shift do you have?
2	FeSCN^{2+} is a mix of an iron/reddish color + thiocyanate has no color	Shifted Right	The Le Chatelier principle w/ reaction from left to right occurs when more reactant was added.
3	$[\text{SCN}^{-}]$ was increased	Shifted Right	As more reactant was added, the concentration of the product goes up
4	Fe^{3+} concentration was changed	Shifted Left	FeCO_3 will form which leads to a lower Fe^{3+} concentration

Lab 1rate And Equilibrium Report Answer

Julia Schneider



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Table of Contents Lab 1rate And Equilibrium Report Answer

1. Understanding the eBook Lab 1rate And Equilibrium Report Answer
 - The Rise of Digital Reading Lab 1rate And Equilibrium Report Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab 1rate And Equilibrium Report Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab 1rate And Equilibrium Report Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab 1rate And Equilibrium Report Answer
 - Personalized Recommendations
 - Lab 1rate And Equilibrium Report Answer User Reviews and Ratings
 - Lab 1rate And Equilibrium Report Answer and Bestseller Lists
5. Accessing Lab 1rate And Equilibrium Report Answer Free and Paid eBooks
 - Lab 1rate And Equilibrium Report Answer Public Domain eBooks
 - Lab 1rate And Equilibrium Report Answer eBook Subscription Services
 - Lab 1rate And Equilibrium Report Answer Budget-Friendly Options
6. Navigating Lab 1rate And Equilibrium Report Answer eBook Formats

- ePub, PDF, MOBI, and More
- Lab 1rate And Equilibrium Report Answer Compatibility with Devices
- Lab 1rate And Equilibrium Report Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab 1rate And Equilibrium Report Answer
 - Highlighting and Note-Taking Lab 1rate And Equilibrium Report Answer
 - Interactive Elements Lab 1rate And Equilibrium Report Answer
- 8. Staying Engaged with Lab 1rate And Equilibrium Report Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab 1rate And Equilibrium Report Answer
- 9. Balancing eBooks and Physical Books Lab 1rate And Equilibrium Report Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab 1rate And Equilibrium Report Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Lab 1rate And Equilibrium Report Answer
 - Setting Reading Goals Lab 1rate And Equilibrium Report Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Lab 1rate And Equilibrium Report Answer
 - Fact-Checking eBook Content of Lab 1rate And Equilibrium Report Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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