

*Lab (25 pts):
Empirical Formula
of Magnesium Oxide*

Name _____
Partner _____
Period _____ Date _____

Purpose: To experimentally determine the empirical formula for magnesium oxide, Mg_xO_y .

Introduction:

By now you should know that a **molecular formula** is a group of elemental symbols and subscript numbers, which represent the composition of a molecule. The molecular formula shows you how many of each atom can be found in a certain molecule. Some examples of molecular formulas, that you are most likely familiar with, are shown in Table A.

Table A - Some Molecular Formulas and the Information they Convey	
Formula	Composition
H_2O	Total of 3 atoms in each molecule - 1 oxygen and 2 hydrogen
O_2	Total of 2 atoms in each molecule - 2 oxygen atoms
CO_2	Total of 3 atoms in each molecule - 1 carbon and 2 oxygen
$\text{C}_6\text{H}_{12}\text{O}_6$	Total of 24 atoms in each molecule - 6 carbon, 12 hydrogen, and 6 oxygen.

While the molecular formula shows the molecule as it actually exists, the **empirical formula** for a compound shows the simplest whole number ratio for the elements in the compounds, like reducing a fraction. To find the empirical formula for a compound, find the largest whole number that can be divided into each of the subscripts evenly (greatest common factor). If no number greater than 1 will go into any of the subscripts evenly, then the empirical formula and the molecular formula for the compound are the same. Table B shows examples of some empirical and molecular formulas.

Table B - Comparing Empirical and Molecular Formulas			
Compound	Molecular Formula	GCF	Empirical Formula
Water	H_2O	1	H_2O
Hydrogen Peroxide	H_2O_2	2	HO
Glucose	$\text{C}_6\text{H}_{12}\text{O}_6$	6	CH_2O
Methane	CH_4	1	CH_4
Ethane	C_2H_6	2	CH_3
Octane	C_8H_{18}	2	C_4H_9

In this lab, you will perform the following reaction:



where you will determine the subscripts, x and y . The mass of the solid (magnesium) will increase as the oxygen gas combines with magnesium metal to form magnesium oxide. You will measure the mass of Mg before the reaction and the mass of Mg_xO_y after the reaction to determine the mass of O that reacted from the air, and the mass of Mg and O in magnesium oxide. Then you will calculate magnesium oxide's empirical formula.

Procedure: Goggles are **ABSOLUTELY** required!! Don't touch a **HOT** crucible--it will burn you!

- 1) Obtain a stainless steel crucible and lid, and record the mass of the crucible and lid together (watch units & sig figs). Slight discoloration of the crucible is normal and will not affect your results. We are using special balanced for this lab, which measure mass to 0.001 g. *Record all masses to 0.001 g (3 decimal pts).* Use gloves or a crucible tong to handle the crucible. When carrying the crucible hold it firmly with crucible tongs, and hold a piece of metal gauze beneath the crucible in case it slips.



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Jo Allan Beran



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Table of Contents Lab Determining An Empirical Formula Mgo Answers

1. Understanding the eBook Lab Determining An Empirical Formula Mgo Answers
 - The Rise of Digital Reading Lab Determining An Empirical Formula Mgo Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Determining An Empirical Formula Mgo Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Determining An Empirical Formula Mgo Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Determining An Empirical Formula Mgo Answers

- Personalized Recommendations
 - Lab Determining An Empirical Formula Mgo Answers User Reviews and Ratings
 - Lab Determining An Empirical Formula Mgo Answers and Bestseller Lists
5. Accessing Lab Determining An Empirical Formula Mgo Answers Free and Paid eBooks
 - Lab Determining An Empirical Formula Mgo Answers Public Domain eBooks
 - Lab Determining An Empirical Formula Mgo Answers eBook Subscription Services
 - Lab Determining An Empirical Formula Mgo Answers Budget-Friendly Options
 6. Navigating Lab Determining An Empirical Formula Mgo Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab Determining An Empirical Formula Mgo Answers Compatibility with Devices
 - Lab Determining An Empirical Formula Mgo Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab Determining An Empirical Formula Mgo Answers
 - Highlighting and Note-Taking Lab Determining An Empirical Formula Mgo Answers
 - Interactive Elements Lab Determining An Empirical Formula Mgo Answers
 8. Staying Engaged with Lab Determining An Empirical Formula Mgo Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Determining An Empirical Formula Mgo Answers
 9. Balancing eBooks and Physical Books Lab Determining An Empirical Formula Mgo Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Determining An Empirical Formula Mgo Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Lab Determining An Empirical Formula Mgo Answers
 - Setting Reading Goals Lab Determining An Empirical Formula Mgo Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Lab Determining An Empirical Formula Mgo Answers

- Fact-Checking eBook Content of Lab Determining An Empirical Formula Mgo Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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