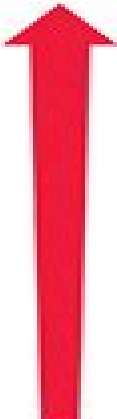
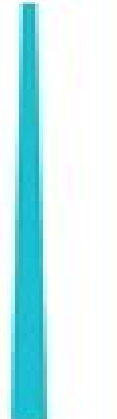
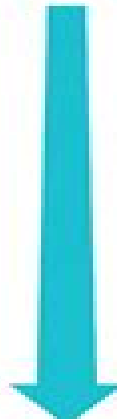


Acids and Bases

MERCK

Acid				Base		
Strength of Acid	Ka	Name	Formula	Formula	Name	Strength of Base
Strongest 	Large	Perchloric acid	HClO_4	ClO_4^-	Perchlorate ion	Weakest 
	3.2×10^8	Hydroiodic acid		I^-	Iodide	
	1.0×10^9	Hydrobromic acid	HBr	Br^-	Bromide	
	1.3×10^4	Hydrochloric acid	HCl	Cl^-	Chloride	
	1.0×10^3	Sulfuric acid	H_2SO_4	HSO_4^-	Hydrogen sulfate ion	
	2.4×10^1	Nitric acid	HNO_3	NO_3^-	Nitrate ion	
	-----	Hydronium ion	H_3O^+	H_2O	Water	
	5.4×10^{-2}	Oxalic acid	$\text{HO}_2\text{C}_2\text{O}_2\text{H}$	$\text{HO}_2\text{C}_2\text{O}_2^-$	Hydrogen oxalate ion	
	1.3×10^{-2}	Sulfurous acid	H_2SO_3	HSO_3^-	Hydrogen sulfite ion	
	1.0×10^{-2}	Hydrogen sulfate ion	HSO_4^-	SO_4^{2-}	Sulfate ion	
Medium 	7.1×10^{-3}	Phosphoric acid	H_3PO_4	H_2PO_4^-	Dihydrogen phosphate ion	Medium 
	7.2×10^{-4}	Nitrous acid	HNO_2	NO_2^-	Nitrite ion	
	6.6×10^{-4}	Hydrofluoric acid	HF	F^-	Fluoride ion	
	1.8×10^{-4}	Methanoic acid	HCO_2H	HCO_2^-	Methanoate ion	
	6.3×10^{-5}	Benzoic acid	$\text{C}_6\text{H}_5\text{COOH}$	$\text{C}_6\text{H}_5\text{COO}^-$	Benzoate ion	
	5.4×10^{-5}	Hydrogen oxalate ion	$\text{HO}_2\text{C}_2\text{O}_2^-$	$\text{O}_2\text{C}_2\text{O}_2^{2-}$	Oxalate ion	
	1.8×10^{-5}	Ethanoic acid	CH_3COOH	CH_3COO^-	Ethanoate (acetate) ion	
	4.4×10^{-7}	Carbonic acid	H_2CO_3	HCO_3^-	Hydrogen carbonate ion	
	1.1×10^{-7}	Hydrosulfuric acid	H_2S	HS^-	Hydrogen sulfide ion	
	6.3×10^{-8}	Dihydrogen phosphate ion	H_2PO_4^-	HPO_4^{2-}	Hydrogen phosphate ion	
Weakest	6.2×10^{-8}	Hydrogen sulfite ion	HS^-	SO_3^{2-}	Sulfite ion	Strongest
	2.9×10^{-8}	Hypochlorous acid	HClO	ClO^-	Hypochlorite ion	
	6.2×10^{-10}	Hydrocyanic acid	HCN	CN^-	Cyanide ion	
	5.8×10^{-10}	Ammonium ion	NH_4^+	NH_3	Ammonia	
	5.8×10^{-10}	Boric acid	H_3BO_3	H_2BO_3^-	Dihydrogen borate	
	4.7×10^{-11}	Hydrogen carbonate ion	HCO_3^-	CO_3^{2-}	Carbonate ion	
	4.2×10^{-13}	Hydrogen phosphate ion	HPO_4^{2-}	PO_4^{3-}	Phosphate ion	
	1.8×10^{-13}	Dihydrogen borate ion	H_2BO_3^-	HBO_3^{2-}	Hydrogen borate ion	
	1.3×10^{-13}	Hydrogen sulfide ion	HS^-	S^{2-}	Sulfide ion	
	1.6×10^{-14}	Hydrogen borate ion	HBO_3^{2-}	BO_3^{3-}	Borate ion	
	-----	water	H_2O	OH^-	Hydroxide	



Lab 16 1 Creating Acid Base Strength Table

Tobias Bleicker

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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