

Experiment V: Equilibrium Constant Lab Report

I. Calibration Curve

- Record the absorbance for each of the standard solutions in the table below. Calculate the concentration of $\text{Fe}(\text{SCN})^{2+}$ in each of the standard solutions and record the values in the table below. Show a sample calculation (for A2) of your work in the space below the table.

Solution	Absorbance	$[\text{Fe}(\text{SCN})^{2+}]$
0mL A-1	0.000 A	0
5mL A-2	0.382 A	1.0×10^{-4}
7mL A-3	0.487 A	1.4×10^{-4}
9mL A-4	0.688 A	1.8×10^{-4}
11mL A-5	0.854 A	2.2×10^{-4}
13mL A-6	1.023 A	2.6×10^{-4}
15mL A-7	1.152 A	3.0×10^{-4}

$$A-2 = \frac{5\text{mL}(0.002\text{M})}{100\text{mL}} = 1.0 \times 10^{-4}$$

$$A-7 = \frac{15\text{mL}(0.002\text{M})}{100\text{mL}} = 3.0 \times 10^{-4}$$

$$A-3 = \frac{7\text{mL}(0.002\text{M})}{100\text{mL}} = 1.4 \times 10^{-4}$$

$$A-4 = \frac{9\text{mL}(0.002\text{M})}{100\text{mL}} = 1.8 \times 10^{-4}$$

$$A-5 = \frac{11\text{mL}(0.002\text{M})}{100\text{mL}} = 2.2 \times 10^{-4}$$

$$A-6 = \frac{13\text{mL}(0.002\text{M})}{100\text{mL}} = 2.6 \times 10^{-4}$$

- Graph $[\text{Fe}(\text{SCN})^{2+}]$ vs. Absorbance to generate your calibration curve for this experiment using Excel. Ensure that all axes are properly labeled and display the equation for the line on the graph. Attach this graph to lab report. Also write out the equation for the calibration curve line below.

Lab 11 Rate And Equilibrium Report Answer

Richard Bailey



Lab 11 Rate And Equilibrium Report Answer:

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