

Date _____ Student Name Victor Carr
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The Identity of an Insoluble Precipitate

Pre-laboratory Assignment

1. The following reactions, shown in unbalanced equations, are pertinent to this experiment. Balance the equations:



2. When a solution containing 1.4 g of $\text{Ba}(\text{NO}_3)_2$ and 2.4 g of $\text{NH}_2\text{SO}_3\text{H}$ is boiled, a precipitate forms. This precipitate may be $\text{Ba}(\text{NH}_2\text{SO}_3)_2$, BaSO_4 , or $\text{Ba}(\text{NH}_2)_2$. You can find these products in the reactions whose equations you balanced in the first question.

- a. Calculate the number of moles of each reactant.

moles of Ba = $\frac{1.4 \text{ g}}{261.34 \text{ g/mol}} = 0.0054 \text{ moles}$?

moles of $\text{NH}_2\text{SO}_3\text{H}$ = $\frac{2.4 \text{ g}}{147.09 \text{ g/mol}} = 0.016 \text{ moles}$?

Identity Of An Insoluble Precipitate Prelab Answers

DP Hallahan



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