

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



### NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride  $\text{SF}_6$
- b. Lithium nitride  $\text{Li}_3\text{N}$
- c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride  $\text{SnCl}_2$
- e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- g. Potassium bromate  $\text{KBrO}_3$
- h. Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- i. Chloric acid  $\text{HClO}_3$
- j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds.

- a.  $\text{CuSO}_4$  Copper (II) sulfate
- b.  $\text{AlF}_3$  aluminum fluoride
- c.  $\text{HI}$  hydrogen iodide
- d.  $\text{NO}$  nitrogen monoxide
- e.  $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- f.  $\text{HNO}_3$  nitric acid
- g.  $\text{HNO}_2$  nitrous acid
- h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \rightarrow 20 \rightarrow 10 \rightarrow 5 \rightarrow 2.5 \rightarrow 1.25$$

1 2 3 4 5 61.5yr

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$100 \rightarrow 50 \rightarrow 25 \rightarrow 12.5\text{mg}$$

1 2 3 12.5mg

3 half lives

### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ \hline & + & \\ & 63.62 & \rightarrow 64 \end{array}$$

# K12 Chemistry A Laboratory Guide Answers

**LP Steffe**



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**The Complete Home Learning Sourcebook** Rebecca Rupp, 1998 Lists all the resources needed to create a balanced curriculum for homeschooling from preschool to high school level

*Nuclear Science Abstracts* ,1976-03     *Illinois Chemistry Teacher* ,2007-09     *Chemistry of Cement* ,1962

*ENC Focus* ,2001     **Resources in Education** ,1999-04     **The Science Teacher** ,2007     **New Horizons in Mathematics and Science Education** ,2001     **Scientific and Technical Aerospace Reports** ,1975     *Internet Directory for Teachers* Grace Jasmine, Julia Jasmine, 1997 These pages guide teachers to the best and most useful resources on the Internet for use in the classroom for career development and for joining the educational community The multi platform CD ROM includes an Internet service provider a search engine and links to sites

**The Wizard of Food Presents 21st Century Reference Guide to Cooking Secrets & Helpful Household Hints** Myles Bader, 2000     **Achieving Indigenous Student Success** Pamela Rose Toulouse, 2017-01-12 In *Achieving Indigenous Student Success* author Pamela Toulouse provides strategies lessons and hands on activities that support both Indigenous and non Indigenous learners in the secondary classroom Read chapters on topics such as Indigenous Pedagogy and Classrooms Considerations Indigenous Self Esteem and Mental Health Activities Differentiated Instruction and Bloom's Taxonomy Attrition Retention Transition and Graduation Continuum Indigenous Themes and Material Resources Culturally Appropriate Secondary Lesson Plans by Subject including English Math Science History Geography Health Physical Education Drama Music Visual Arts Technological Studies Business Studies Indigenous Worldviews Guidance and Career Studies and Social Studies and the

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