

## Energy and Chemical Change

### Section 15.1 Energy

pages 516–522

#### Practice Problems

pages 519–521

1. A fruit and oatmeal bar contains 142 nutritional Calories. Convert this energy to calories.

$$142 \text{ Calories} = 142 \text{ kcal}$$

$$142 \text{ kcal} \times \frac{1000 \text{ cal}}{1 \text{ kcal}} = 142,000 \text{ cal}$$

2. An exothermic reaction releases 86.5 kJ. How many kilocalories of energy are released?

$$86.5 \text{ kJ} \times \frac{1 \text{ kcal}}{4.184 \text{ kJ}} = 20.7 \text{ kcal}$$

3. **Challenge** Define a new energy unit, named after yourself, with a magnitude of one-tenth of a calorie. What conversion factors relate this new unit to joules? To Calories?

$$\text{Unit X} = 0.1 \text{ cal}$$

$$1 \text{ cal} = 4.184 \text{ J}$$

$$\text{X} = (0.1 \text{ cal})(4.184 \text{ J/cal}) = 0.4184 \text{ J}$$

$$1 \text{ cal} = 0.001 \text{ Calorie}$$

$$\text{X} = (0.1 \text{ cal})(1 \text{ Cal}/1000 \text{ cal}) = 0.0001 \text{ Calorie}$$

4. If the temperature of 34.4 g of ethanol increases from 25.0°C to 78.8°C, how much heat has been absorbed by the ethanol? Refer to Table 15.2.

$$q = c \times m \times \Delta T$$

$$q = 2.44 \text{ J/(g} \cdot ^\circ\text{C)} \times 34.4 \text{ g} \times 53.8^\circ\text{C} = 4.52 \times 10^3 \text{ J}$$

5. A 155-g sample of an unknown substance was heated from 25.0°C to 40.0°C. In the process, the substance absorbed 5696 J of energy. What is the specific heat of the substance? Identify the substance among those listed in Table 15.2 on page 520.

$$q = c \times m \times \Delta T$$

$$c = \frac{q}{m\Delta T} = \frac{(5696 \text{ J})}{(155 \text{ g})(40.0 - 25.0^\circ\text{C})} = 2.45 \text{ J/(g} \cdot ^\circ\text{C)}$$

The specific heat is very close to the value for ethanol.

6. **Challenge** A 4.50-g nugget of pure gold absorbed 276 J of heat. The initial temperature was 25.0°C. What was the final temperature?

$$q = c \times m \times \Delta T$$

$$\Delta T = \frac{q}{cm} = \frac{(276 \text{ J})}{(0.129 \text{ J/g} \cdot ^\circ\text{C})(4.50 \text{ g})} = 475^\circ\text{C}$$

$$\Delta T = T_f - T_i = T_f - 25.0^\circ\text{C} = 475^\circ\text{C}$$

$$T_f = 5.00 \times 10^2^\circ\text{C}$$

### Section 15.1 Assessment

page 522

7. **Explain** how energy changes from one form to another in an exothermic reaction. In an endothermic reaction.

Chemical potential energy changes to heat in exothermic reactions and the heat is released. In endothermic reactions, heat is absorbed and changed to chemical potential energy.

8. **Distinguish** between kinetic and potential energy in the following examples: two separated magnets; an avalanche of snow; books on library shelves; a mountain stream; a stock-car race; separation of charge in a battery.

Two separated magnets illustrate potential energy. In a snow avalanche, positional potential energy is changing to kinetic energy. Books on a shelf illustrate positional potential energy. As water races down a mountain stream, positional potential energy is changing to kinetic energy. In a stock-car race, chemical potential energy is being changed to kinetic energy. The separation of charge in a battery illustrates electrical potential energy.

9. **Explain** how the light and heat of a burning candle are related to chemical potential energy.

Chemical potential energy, contained in the candle, is changed to energy in the form of light and heat and released as the chemical combustion reaction takes place.

# Energy And Chemical Change Solutions Manual

**Joseph Albert Babor, Alexander  
Lehrman**



## **Energy And Chemical Change Solutions Manual:**

**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03

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A. Trapp, Peter Atkins, Julio dePaula, 2009-12-18. With its modern emphasis on the molecular view of physical chemistry, its wealth of contemporary applications, vivid full-color presentation, and dynamic new media tools, the thoroughly revised new edition is again the most modern, most effective full-length textbook available for the physical chemistry classroom. Available in Split Volumes. For maximum flexibility in your physical chemistry course, this text is now offered as a traditional text or in two volumes: Volume 1: Thermodynamics and Kinetics, ISBN 1 4292 3127 0; Volume 2: Quantum Chemistry, Spectroscopy, and Statistical Thermodynamics, ISBN 1 4292 3126 2.

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Department, **Solutions Manual to Accompany Elements of Physical Chemistry** C. A. Trapp, M. P. Cady, 2007. The Solutions manual to accompany Elements of Physical Chemistry, 4e, contains full worked solutions to all end-of-chapter exercises featured in the book.

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Lehrman, 1929. **Solutions Manual to Accompany Elements of Physical Chemistry** David Smith, 2013-05-30. The Solutions Manual to accompany Elements of Physical Chemistry, 6th edition, contains full worked solutions to all end-of-chapter discussion questions and exercises featured in the book. The manual provides helpful comments and friendly advice to aid understanding. It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured

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Laboratory Manual of Inorganic Chemistry for Colleges Lyman Churchill Newell, 1916

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**Solutions Manual for Principles of Physical Chemistry, 3rd Edition** Hans Kuhn, David H. Waldeck, Horst-Dieter Försterling, 2024-10-29 This is a Solutions Manual to accompany with solutions to the exercises in the main volume of Principles of Physical Chemistry Third Edition This book provides a unique approach to introduce undergraduate students to the concepts and methods of physical chemistry which are the foundational principles of Chemistry The book introduces the student to the principles underlying the essential sub fields of quantum mechanics atomic and molecular structure atomic and molecular spectroscopy statistical thermodynamics classical thermodynamics solutions and equilibria electrochemistry kinetics and reaction dynamics macromolecules and organized molecular assemblies Importantly the book develops and applies these principles to supramolecular assemblies and supramolecular machines with many examples from biology and nanoscience In this way the book helps the student to see the frontier of modern physical chemistry developments The book begins with a discussion of wave particle duality and proceeds systematically to more complex chemical systems in order to relate the story of physical chemistry in an intellectually coherent manner The topics are organized to correspond with those typically given in each of a two course semester sequence The first 13 chapters present quantum mechanics and spectroscopy to describe and predict the structure of matter atoms molecules and solids Chapters 14 to 29 present statistical thermodynamics and kinetics and applies their principles to understanding equilibria chemical transformations macromolecular properties and supramolecular machines Each chapter of the book begins with a simplified view of a topic and evolves to more rigorous description in order to provide the student and instructor flexibility to choose the level of rigor and detail that suits them best The textbook treats important new directions in physical chemistry research including chapters on macromolecules principles of interfaces and films for organizing matter and supramolecular machines as well as including discussions of modern nanoscience spectroscopy and reaction dynamics throughout the text

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*Student Solutions Manual: Ssm Chemistry* Deborah Wiegand, 2003 This manual contains complete worked out solutions to all follow up problems and about half of all the chapter problems Each chapter of solutions opens with a summary of the text chapter content and a list of key equations needed to solve the problems

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web mar 2 2004 this movie lines up nicely with the book number the stars by lois lowry it is the incredible true story of how the danish people helped to save the lives of their fellow jewish neighbors and friends from the nazi invasion i ve watched it 20 times and still tear up every time read the book too it is not exactly the same but similar enough

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web september 5 2017 i posted this video late saturday night september 2nd after making a surprise announcement at my san francisco comic con panel i was tired but felt duty bound to explain my announcement the true purpose of the video was to see if there was genuine support to have number the stars adapted for the screen

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web october 18 2017 3 minutes lois lowry s 1989 novel number the stars didn t just win a prestigious newbery medal and make you cry as you wrote your book report it was also based on a true story as lowry told the reading teacher in 1990

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web number the stars is an upcoming 2019 french german swedish danish american animated historical fiction adventure fantasy film based on the novel of the same name this film is about the escape of a jewish family from copenhagen during the second world war it is directed by adam katz and produced by lois lowry and john davis it is set to

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web number the stars is a historical fiction novel by lois lowry about the escape of a family of jews from copenhagen denmark during world war ii the story follows ten year old annemarie johansen who helps her best friend ellen rosen who is pretending to be her sister by smuggling jews to sweden the novel was awarded the newbery medal in

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web nov 22 2020 is there a movie for the novel number the stars this movie pairs well with the book number the stars by lois lowry it is an age appropriate movie for kids in upper elementary who are wanting to learn more about events of ww2 this disney made for tv movie was inspired by the people of denmark who work together to save their jewish

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