

## Empirical Formulas Practice Packet

Show all Calculations. Write final answers in the correct number of significant figures and with the correct units

1. A 7.546 g sample of an unknown compound was analyzed and found to contain only nitrogen and oxygen. If the mass of nitrogen present in the sample was 1.957 g, determine the empirical formula of the compound.
2. Serotonin is a compound that conducts nerve impulses in the brain. Serotonin has the following percent composition by mass: 68.2% carbon, 6.86% hydrogen, 15.9% nitrogen, and 9.08% oxygen.
  - a. What is its empirical formula?
  - b. What is its molecular formula if its molar mass is about 176 g/mol?
3. The koala dines exclusively on eucalyptus leaves. Its digestive system detoxifies the eucalyptus oil that can poison other animals. The chief constituent in eucalyptus oil is a substance called eucalyptol, which has the following mass percent composition: 77.87% carbon, 11.76% hydrogen, and the remainder is oxygen.
  - a. What is its empirical formula?
  - b. What is its molecular formula if its molar mass is about 154 g/mol?
4. Determine the molecular formulas for the following compounds given their empirical formulas and their molar masses: a) isoamyl formate: empirical formula= $\text{C}_3\text{H}_6\text{O}$ , molar mass=116 g/mol; b) lactic acid: empirical formula= $\text{CH}_2\text{O}$ , molar mass=90.1 g/mol; and c) opionic acid: empirical formula= $\text{C}_3\text{H}_2\text{O}$ , molar mass=210.2 g/mol.

# Empirical Formula Packet

**Patrick Vollmar**



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