

Empirical Formula and Molecular Formula

Lab 10: Molecular Formulas and Empirical Formulas

Lab 10

Students will learn how to determine the empirical formula of a compound from its mass percent composition. They will also learn how to determine the molecular formula of a compound from its mass percent composition and its molar mass. This activity will require the use of a calculator.

Model 1: Comparison of Percent Composition and Empirical Formula

Compound	Empirical Formula	Molecular Formula	Empirical Formula	Percent Composition (Carbon)	Percent Composition (Hydrogen)
1	ethane	C_2H_6	CH_3		
2	propane	C_3H_8	$CH_2.67$		
3	ethylene (acetylene)	C_2H_2	CH		
4	benzene	C_6H_6	CH		

1. Looking at the table, what are some similarities between the percent composition and empirical formula data in Model 1. What are some differences? (You may calculate if you wish.)
2. Identify the compounds in Model 1 that have the same empirical formula.
3. Identify the compounds in Model 1 that have the same percent composition.
4. Using the compounds in Model 1, suggest an explanation as to why compounds with different molecular formulas but the same empirical formula have the same percent composition.
5. For each substance in Model 1, give evidence that the molecular formula and empirical formula are exactly the same. What mathematical operation did you use to determine the molecular formula? (You may use the compound's molecular formula.)

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